

1. Answer the following in the space provided. Where required, be sure to use correct scientific notation, significant digits and including proper units of measure. (12 marks)

(a) the product(s) of the reaction of $SO_3(g) + H_2O(l) \rightarrow$

(b) the product(s) of the reaction of $Ba(s) + H_2O(l) \rightarrow$

(c) the product(s) of the reaction of $Na_2CO_3(s) + HNO_3(aq) \rightarrow$

(d) the product(s) of the reaction of $Ag(s) + Pb(NO_3)_2(aq) \rightarrow$

(e) the number of moles in 9.97×10^{25} molecules of CO_2

(f) the number of H atoms in 2.02 moles of $CH_3(CH_2)_8COOH$

(g) the number of moles in a 71.92 g sample of $MgSO_4$

(h) the mass of 3.33×10^{20} molecules of chlorine

(i) percentage composition
of ammonium phosphate

chemical formula

molar mass

percentage composition

$H_2SO_4(aq)$	
$Ba(OH)_2(aq) + H_2(g)$	
$NaNO_3(aq) + H_2O(l) + CO_2(g)$	
no rxn.	
166 mol	
2.43×10^{25} H atoms	
0.5974 mol	
0.0392 g	
$(NH_4)_3PO_4$	
149.12 $\frac{g}{mol}$	
28.19% N	8.13% H
20.77% P	42.92% O

Answer the following with complete solutions, be sure to indicate all units of measure and express final answer to correct significant digits.

2. A molecular compound containing the unknown element X has the formula X_4O_{10} . If it is known that the compound is 56.36% oxygen by mass, calculate the molar mass of element X and state its identity. (2 marks)

$$56.36\% = \frac{10(16.00 \frac{g}{mol})}{M_{X_4O_{10}}} \times 100\% \quad 283.9 \frac{g}{mol} = 4(M_X) + 10(16.00 \frac{g}{mol})$$

$$M_X = 30.97 \frac{g}{mol}$$

$$M_{X_4O_{10}} = 283.9 \frac{g}{mol}$$

$\therefore X$ is phosphorus (P)

Answer Key For Chemical Quantities Test

JS Bruner



Answer Key For Chemical Quantities Test:

CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

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Air Force Manual United States. Department of the Air Force, 1973 *Chemistry: The Central Science* Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation **The Chemical Gazette, Or, Journal of Practical Chemistry, in All Its Applications to**

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