

PRACTICE!

LAP SCIENCE



- Describe one example of how climate affects people.
- Explain the difference between anecdotal and scientific evidence. Give one example (can be made up) of each for climate change.
- What does albedo mean?
- Compare the albedo of an area that is covered in ice to an area that is covered with dark soil.
- Explain what is meant by the term "Earth's net radiation budget".
- Distinguish between the angle of inclination and the angle of incidence.
- Compare and contrast conduction and convection.
- Describe the Coriolis effect. Which direction does it make currents go in hemisphere?

THERMAL ENERGY TRANSFER IN THE BIOSPHERE

Thermal energy is transferred through the hydrosphere from the equator to the poles. Polar regions consistently are the coldest.

SPECIFIC HEAT CAPACITY: the amount of thermal energy required to raise the temperature of 1 gram of a substance by 1°C.

QUANTITY OF THERMAL ENERGY:

formula

$$Q = mc\Delta T$$

EXAMPLES

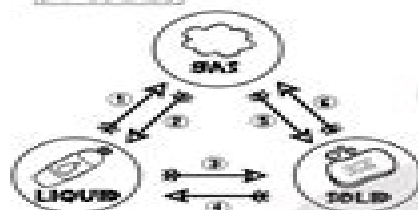
1. A 350-g mass of water at 12.0°C is allowed to warm up to 30.0°C. Determine the amount of thermal energy, Q , absorbed. The theoretical specific heat capacity of water is 4.19 J/g°C.

2. A 100-g mass of iron is at 30°C. Determine the final temperature of the iron after it absorbs 56 J of thermal energy. The theoretical specific heat capacity of iron is 0.449 J/g°C.

usually, when you add energy to a substance, what happens to its temperature?

but at two specific temperatures, whatever the amount of energy added, the temperature remains constant. What happens to the energy then?

LAP SCIENCE



- 1
- 2
- 3
- 4
- 5
- 6

during a phase change, the temperature remains constant.

	HEAT OF FUSION	HEAT OF VAPORIZATION
The amount of energy released or absorbed when a substance changes between a...		
Formula		

$$H_{\text{fusion}} = Q = m \cdot h_{\text{fusion}}$$

$$H_{\text{vapor}} = Q = m \cdot h_{\text{vapor}}$$

reminder: moles

formula

$$n = \frac{m}{M}$$

THERMAL ENERGY TRANSFER IN THE BIOSPHERE

Specific heat capacity is the amount of thermal energy required to raise the temperature of 1 gram of a substance by 1°C.

Specific heat capacity of different substances is different. Water has a high specific heat capacity.

formula

$$Q = mc\Delta T$$

Q = thermal energy (J)
m = mass (g)
c = specific heat capacity (J/g°C)
 ΔT = change in temperature (°C)

examples

$$Q = mc\Delta T$$

Q = 100 J
m = 100 g
c = 1 J/g°C
 ΔT = 1°C

1. A 100-g mass of water at 12.0°C is allowed to warm up to 30.0°C. Determine the amount of thermal energy, Q , absorbed. The theoretical specific heat capacity of water is 4.19 J/g°C.

2. A 100-g mass of iron is at 30°C. Determine the final temperature of the iron after it absorbs 56 J of thermal energy. The theoretical specific heat capacity of iron is 0.449 J/g°C.

usually, when you add energy to a substance, what happens to its temperature?

IT gets hotter

but at two specific temperatures, whatever the amount of energy added, the temperature remains constant. What happens to the energy then?

it goes into a phase change

during a phase change, the temperature remains constant.

1: evaporation
2: condensation
3: freezing
4: melting
5: sublimation
6: deposition

	HEAT OF FUSION	HEAT OF VAPORIZATION
The amount of energy released or absorbed when a substance changes between a...		
Formula		

reminder: moles

formula

$$n = \frac{m}{M}$$

Answer Key For Science 10 Alberta

Mitsuru Ishizuka, Abdul Satter



Answer Key For Science 10 Alberta:

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Table of Contents Answer Key For Science 10 Alberta

1. Understanding the eBook Answer Key For Science 10 Alberta
 - The Rise of Digital Reading Answer Key For Science 10 Alberta
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Science 10 Alberta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Science 10 Alberta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Science 10 Alberta
 - Personalized Recommendations
 - Answer Key For Science 10 Alberta User Reviews and Ratings
 - Answer Key For Science 10 Alberta and Bestseller Lists

5. Accessing Answer Key For Science 10 Alberta Free and Paid eBooks
 - Answer Key For Science 10 Alberta Public Domain eBooks
 - Answer Key For Science 10 Alberta eBook Subscription Services
 - Answer Key For Science 10 Alberta Budget-Friendly Options
6. Navigating Answer Key For Science 10 Alberta eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For Science 10 Alberta Compatibility with Devices
 - Answer Key For Science 10 Alberta Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Science 10 Alberta
 - Highlighting and Note-Taking Answer Key For Science 10 Alberta
 - Interactive Elements Answer Key For Science 10 Alberta
8. Staying Engaged with Answer Key For Science 10 Alberta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Science 10 Alberta
9. Balancing eBooks and Physical Books Answer Key For Science 10 Alberta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Science 10 Alberta
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key For Science 10 Alberta
 - Setting Reading Goals Answer Key For Science 10 Alberta
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key For Science 10 Alberta
 - Fact-Checking eBook Content of Answer Key For Science 10 Alberta
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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