

Chapter 3: Stoichiometry

1. The atomic mass of rhenium is 186.2. Given that 37.1% of natural rhenium is rhenium-185, what is the other stable isotope?

A) $^{188}_{75}\text{Re}$
B) $^{187}_{75}\text{Re}$
C) $^{189}_{75}\text{Re}$
D) $^{181}_{75}\text{Re}$
E) $^{190}_{75}\text{Re}$

ANS: B

PTS: 1

DIF: Moderate

REF: 3.2

KEY: Chemistry / general chemistry / early atomic theory / atomic theory of matter / atomic weight / mass spectrometry

MSC: Quantitative

2. Consider the element indium, atomic number 49, atomic mass 114.82 g. The nucleus of an atom of indium-112 contains

A) 49 protons, 63 neutrons, 49 electrons
B) 49 protons, 49 neutrons
C) 49 protons, 49 alpha particles
D) 49 protons, 63 neutrons
E) 49 protons, 112 neutrons

ANS: D

PTS: 1

DIF: Easy

REF: 3.2

KEY: Chemistry / general chemistry / early atomic theory / atomic theory of matter / structure of the atom

MSC: Conceptual

3. A hypothetical element consists of two isotopes of masses 69.95 amu and 71.95 amu with abundances of 25.7% and 74.3%, respectively. What is the average atomic mass of this element?

A) 70.95 amu
B) 69.95 amu
C) 70.5 amu
D) 71.4 amu
E) 71.95 amu

ANS: D

PTS: 1

DIF: Easy

REF: 3.2

KEY: Chemistry / general chemistry / early atomic theory / atomic theory of matter / atomic weight / mass spectrometry

MSC: Quantitative

4. Naturally occurring copper exists in two isotopic forms, ^{63}Cu and ^{65}Cu . The atomic mass of copper is 63.55 amu. What is the approximate natural abundance of ^{63}Cu ?

A) 67%
B) 69%
C) 70%
D) 75%
E) 78%

ANS: C

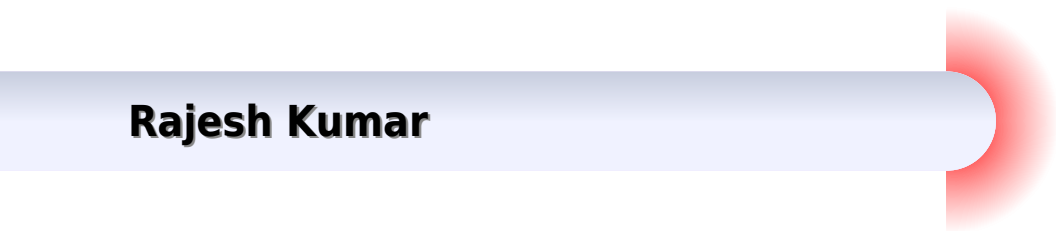
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