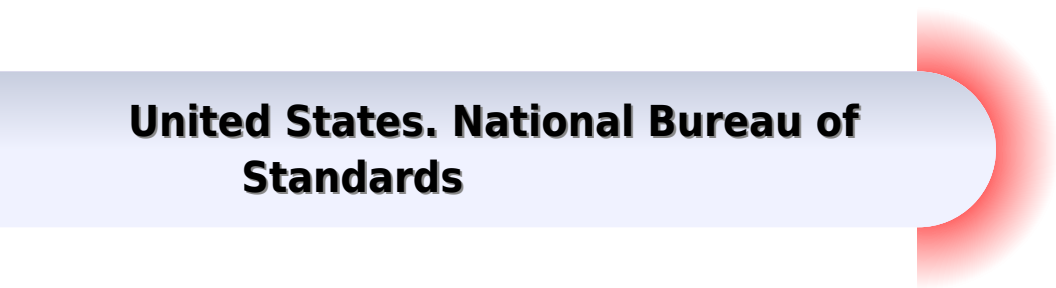


Kinetic Molecular Theory Worksheet

1. The word *kinetic* comes from a Greek word that means "to move." The **kinetic molecular theory** is based upon the assumption that particles of matter (atoms or molecules) are in constant **random motion**.
 2. Of the three states of matter, which one has the most kinetic energy? **Gas**.
 3. Which state of matter has particles that are separated by the largest distance? **Gas**.
- The diagram consists of three rectangular boxes labeled 'Solid', 'Liquid', and 'Gas' below them. The 'Solid' box is filled with a dense, regular grid of blue dots. The 'Liquid' box contains blue dots that are more disordered and less densely packed than in the solid state. The 'Gas' box contains a few blue dots that are widely spaced and scattered throughout the container.
4. A scientific theory is an explanation of some type of natural phenomena. Theories are normally developed from careful study of the way the world behaves. Let's look at how gases behave and see if the kinetic molecular theory makes sense.
 5. Compared to liquids and solids, gases tend to have **lower** densities. This can be explained because the particles of gas are **in constant motion**.
 6. If you apply pressure to a sample of gas, it is fairly easy to **compress** its volume (think about what would happen to a balloon if you squeeze it gently). This can be explained because there is a lot of **empty space** between gas particles.
 7. If someone sprays perfume in one corner of the room, eventually a person on the other side of the room can smell it. This can be explained because gas particles move **quick** and **random**. In general, we would expect lighter gas particles to travel **faster** than heavier gas particles.
 8. When two gases mix together or move through each other, this process is known as **diffusion**. When gas particles escape out of a tiny hole in a container, this process is known as **effusion**. You should know the difference between these two words so you can avoid any confusion!
 9. Kinetic molecular theory can be summarized as follows:
 - a. Gas particles are in constant **motion**.
 - b. Gas particles are separated by relatively **large** distances.
 - c. When gas particles collide, they **do not transfer** kinetic energy.
 - d. Gas particles have **no** attractive or repulsive forces between them.
 - e. The kinetic energy of a gas is dependent on the **temperature** of the gas.

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