

Unit 3 Worksheet 4 – Quantitative Energy Problems Part 2

Energy constants (H₂O)

334 J/g	Heat of fusion (melting or freezing) H_f
2260 J/g	Heat of vaporization (evaporating or condensing) H_v
2.1 J/g°C	Heat capacity (c) of solid water
4.18 J/g°C	Heat capacity (c) of liquid water

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

1. How much energy must be absorbed by a 150 g sample of ice at 0.0 °C that melts and then warms to 25.0°C?
2. Suppose in the Icy Hot lab that the burner transfers 325 kJ of energy to 450 g of liquid water at 20.°C. What mass of the water would be boiled away?
3. A 12oz can of soft drink (assume $m = 340$ g) at 25°C is placed in a freezer where the temperature is - 12 °C. How much energy must be removed from the soft drink for it to reach this temperature?

Answer For Quantitative Energy Problems

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