

8 • Electron Configurations & Periodicity

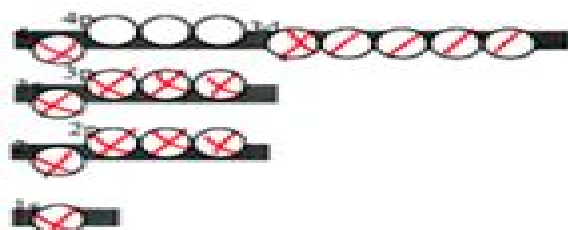
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
 Ar $[Ne] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
 Mg $[Ne] 3s^2$
- N $1s^2 2s^2 2p^3$
 N $[He] 2s^2 2p^3$
- Li $1s^2 2s^1$
 Li $[He] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
 P $[Ne] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
 Cl $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$

Fe $[Ar] 3d^6 4s^2$

Answer Key Electron Configurations Of Atoms And Ions Lab

RJ Shavelson



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Table of Contents Answer Key Electron Configurations Of Atoms And Ions Lab

1. Understanding the eBook Answer Key Electron Configurations Of Atoms And Ions Lab
 - The Rise of Digital Reading Answer Key Electron Configurations Of Atoms And Ions Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Electron Configurations Of Atoms And Ions Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Electron Configurations Of Atoms And Ions Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Electron Configurations Of Atoms And Ions Lab
 - Personalized Recommendations
 - Answer Key Electron Configurations Of Atoms And Ions Lab User Reviews and Ratings
 - Answer Key Electron Configurations Of Atoms And Ions Lab and Bestseller Lists
5. Accessing Answer Key Electron Configurations Of Atoms And Ions Lab Free and Paid eBooks
 - Answer Key Electron Configurations Of Atoms And Ions Lab Public Domain eBooks
 - Answer Key Electron Configurations Of Atoms And Ions Lab eBook Subscription Services
 - Answer Key Electron Configurations Of Atoms And Ions Lab Budget-Friendly Options

6. Navigating Answer Key Electron Configurations Of Atoms And Ions Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Electron Configurations Of Atoms And Ions Lab Compatibility with Devices
 - Answer Key Electron Configurations Of Atoms And Ions Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Electron Configurations Of Atoms And Ions Lab
 - Highlighting and Note-Taking Answer Key Electron Configurations Of Atoms And Ions Lab
 - Interactive Elements Answer Key Electron Configurations Of Atoms And Ions Lab
8. Staying Engaged with Answer Key Electron Configurations Of Atoms And Ions Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Electron Configurations Of Atoms And Ions Lab
9. Balancing eBooks and Physical Books Answer Key Electron Configurations Of Atoms And Ions Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Electron Configurations Of Atoms And Ions Lab
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Electron Configurations Of Atoms And Ions Lab
 - Setting Reading Goals Answer Key Electron Configurations Of Atoms And Ions Lab
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Electron Configurations Of Atoms And Ions Lab
 - Fact-Checking eBook Content of Answer Key Electron Configurations Of Atoms And Ions Lab
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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