

Name: Nicholas Akkanen

Date: October 6 2021

Student Exploration: Diffusion

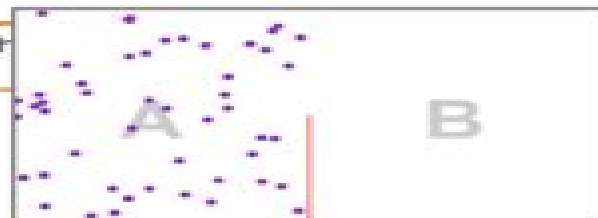
Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: absolute zero, controlled experiment, diffusion, dynamic equilibrium, Kelvin scale, kinetic energy

Prior Knowledge Question (Do this BEFORE using the Gizmo.)

Have you ever smelled microwave popcorn? The oddly enticing scent can fill a whole house. How do you think the smell of popcorn spreads through the air?

This is through the process of diffusion, the gases spread throughout the chamber from high concentration to low concentration.



Gizmo Warm-up

Smells are carried by tiny particles that move through the air. The *Diffusion Gizmo* shows gas particles in a chamber that is divided into two regions by a partial wall. Click **Play** () and observe.

1. Describe the motion of the gas particles.

The gas particles are fast-moving and constantly in motion, bouncing off the walls. In the beginning, many particles move to side B as it has a lower concentration than side A.

2. Over time, what is happening?

Over time, the concentrations of particles even out on both sides (become isotonic), due to the movement of particles from a high concentration to a low concentration.

This process, in which particles move from an area of high concentration (region A) to an area of low concentration (region B), is called **diffusion**.

3. Select the BAR CHART tab, and observe the chart for a few minutes. After the first 30 seconds or so, how much do the numbers of particles in each region change?

The numbers of particles in each region do not change drastically after 30 seconds of net movement. For example, in region B at 0 seconds, the number of particles was 0 and only increased to 10 at 30 seconds.

When the numbers don't change much, the particles are said to be in **dynamic equilibrium**.

Answer Key Exploration Guide Diffusion

Elisabeta Chelsea Wade



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 - The Rise of Digital Reading Answer Key Exploration Guide Diffusion
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Exploration Guide Diffusion
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Exploration Guide Diffusion
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Exploration Guide Diffusion
 - Personalized Recommendations
 - Answer Key Exploration Guide Diffusion User Reviews and Ratings
 - Answer Key Exploration Guide Diffusion and Bestseller Lists

5. Accessing Answer Key Exploration Guide Diffusion Free and Paid eBooks
 - Answer Key Exploration Guide Diffusion Public Domain eBooks
 - Answer Key Exploration Guide Diffusion eBook Subscription Services
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6. Navigating Answer Key Exploration Guide Diffusion eBook Formats
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 - Answer Key Exploration Guide Diffusion Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Exploration Guide Diffusion
 - Highlighting and Note-Taking Answer Key Exploration Guide Diffusion
 - Interactive Elements Answer Key Exploration Guide Diffusion
8. Staying Engaged with Answer Key Exploration Guide Diffusion
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Exploration Guide Diffusion
9. Balancing eBooks and Physical Books Answer Key Exploration Guide Diffusion
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Exploration Guide Diffusion
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Exploration Guide Diffusion
 - Setting Reading Goals Answer Key Exploration Guide Diffusion
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Exploration Guide Diffusion
 - Fact-Checking eBook Content of Answer Key Exploration Guide Diffusion
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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