

CHARLES LAW WORKSHEET #2

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

NAME KRY

1. A gas measures 150 mL at 1.0 atmospheres and 20°C.
Calculate its volume in mL at 0°C and 1.0 atmospheres.

$$V_1 = 150 \text{ mL}$$

$$T_1 = 20^\circ\text{C} + 273 = 293 \text{ K}$$

$$T_2 = 0^\circ\text{C} + 273 = 273 \text{ K}$$

$$\frac{150 \text{ mL}}{293 \text{ K}} = \frac{V_2}{273 \text{ K}}$$

$$\frac{(150 \text{ mL})(273 \text{ K})}{(293 \text{ K})} = \frac{(293 \text{ K})V_2}{(293 \text{ K})}$$

$$V_2 = 140 \text{ mL}$$

2. A gas occupies a volume of 4.50 liters at 27°C. At what temperature in °C would the volume be 6.0 liters, the pressure remaining constant?

$$V_1 = 4.50 \text{ L} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$V_2 = 6.0 \text{ L} \quad T_2 = ?$$

$$\frac{4.50 \text{ L}}{300 \text{ K}} = \frac{6.0 \text{ L}}{T_2}$$

$$\frac{(4.50 \text{ L})T_2}{4.50 \text{ L}} = \frac{(300 \text{ K})(6.0 \text{ L})}{(4.50 \text{ L})}$$

$$T_2 = 400 \text{ K} - 273 \text{ K} = 127^\circ\text{C}$$

3. A sample of gas occupies 220 mL at 10°C and 750 torr.
What volume in mL will the gas have at 20°C and 750 torr?

$$V_1 = 220 \text{ mL} \quad T_1 = 10^\circ\text{C} + 273 = 283 \text{ K}$$

$$V_2 = ? \quad T_2 = 20^\circ\text{C} + 273 = 293 \text{ K}$$

$$\frac{220 \text{ mL}}{283 \text{ K}} = \frac{V_2}{293 \text{ K}}$$

$$\frac{(220 \text{ mL})(293 \text{ K})}{283 \text{ K}} = \frac{(283 \text{ K})V_2}{283 \text{ K}}$$

$$V_2 = 228 \text{ mL}$$

4. A gas occupies a volume of 90.0 mL at 27°C and 740 torr.
What volume in mL will the gas have at 5°C and 740 torr?

$$V_1 = 90.0 \text{ mL} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$T_2 = 5^\circ\text{C} + 273 = 278 \text{ K}$$

$$V_2 = 83.4 \text{ mL}$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \Rightarrow \frac{90.0 \text{ mL}}{300 \text{ K}} = \frac{V_2}{278 \text{ K}}$$

$$\frac{(278 \text{ K})(90.0 \text{ mL})}{(300 \text{ K})} = \frac{V_2(300 \text{ K})}{300 \text{ K}}$$

5. A gas occupies a volume of 120 mL at 27°C and 630 torr.
At what temperature in °C would the volume be 80.0 mL at 630 torr?

$$V_1 = 120 \text{ mL} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$V_2 = 80.0 \text{ mL}$$

$$\frac{(120 \text{ mL})T_2}{120 \text{ mL}} = \frac{(80.0 \text{ mL})(300 \text{ K})}{(120 \text{ mL})}$$

$$\frac{120 \text{ mL}}{300 \text{ K}} = \frac{80.0 \text{ mL}}{T_2}$$

$$T_2 = 200 \text{ K} - 273 \text{ K} = -73^\circ\text{C}$$

Answer Key For Boyles Law

Sabine Zange



Answer Key For Boyles Law:

Reviewing **Answer Key For Boyles Law**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Answer Key For Boyles Law**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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