

13. Complete the following table for the three experiments in Model 2.

	Experiment C	Experiment D	Experiment E
Independent Variable	# of particles	Temp.	Internal External Pressure
Dependent Variable	Volume	Volume	Volume
Controlled Variable(s)	Temp / I+E pressure	External + Internal Pressure / # of particles	Temp / # of particles

14. Provide a molecular level explanation for the increase in volume among the balloons in Experiment C. (How often and/or how hard are the molecules hitting the sides of the container?)

of molecules $\uparrow$, so number of collisions $\uparrow$ + Volume along with it

collisions $\uparrow$ energy stays the same

15. Provide a molecular level explanation for the increase in volume among the balloons in Experiment D.

Temperature $\uparrow$ increasing the kinetic energy of the particles as well
which $\uparrow$ collisions thus $\uparrow$ pressure

16. Provide a molecular level explanation for the increase in volume among the balloons in Experiment E.

if external forces decrease then it allows internal forces $\uparrow$
thus $\uparrow$ volume

Answer Key For Pogil Chemistry Gas Variables

Yi-Tong Ma



Answer Key For Pogil Chemistry Gas Variables:

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