

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bond: **electronegativity**
- In order to understand bond of atoms/molecules we need to know how they are held together, energy
- In order to know a bond between two atoms, energy must be added to separate the atoms

Electronegativity

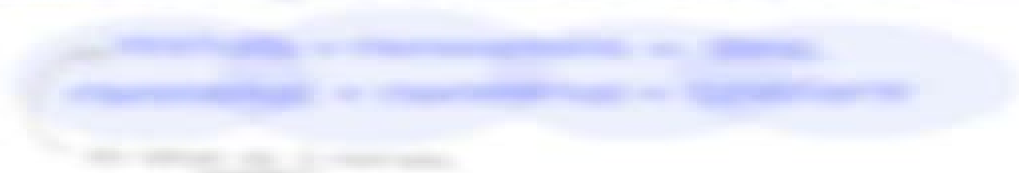
- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent bond
- More electronegative atom is more likely to attract electrons in a bond, making it more negative charge
- Electronegativity increases with the strength of a bond
- From electronegativity values, we can predict the type of bond: **ionic bond**



Covalent bonding

Two atoms share their valence electrons, resulting from bond of electrons from one atom to another atom.

Molecule is a group of atoms held together by chemical bonds.



Electronegativity is the tendency of an atom to attract electrons towards itself in a covalent bond. It is a measure of the ability of an atom to attract electrons.

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