- According to the Bronsted—Lowry definition, which chemical species can function both as an acid and as a base?
 - a. CI = 6. SO₂ =
 - SO/**
 NH/*
 - d. HCO₂
 - e. H₂O*
- 2. In the reaction CN⁻ + H₂O ← → HCN + OH⁻ which is an acid—base conjugate pair?
 - a. H₂O and HCN
 - b. H₂O and OH⁻
 - c. ON and H₂O
 - d. HCN and OH"
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX
 - a. It is less basic than a I M solution of NaY.
 - b. It is more basic than a I M solution of NaY.
 - It yields a neutral solution.
 - It is more concentrated than a 1 M solution of NaY.
- 4. The pH of a 0.03 M HCI solution is
 - a. 1.5
 - 6. 25
 - c. 3.5
 - d. 12.5
- The pH of a 1.0 x 10⁻³ M Ba(OH)₂ solution at 25°C is
 - a. 2.7
 - b. 3.0
 - c. 11.0
 - d. 11.3
- 6. Which is the strongest acid?
 - a. HCIO
 - A. HCIO₂
 - c. HOIO₃
 - d. HClO₄
- 7. Which salt reacts with water (hydrolyzes) to produce a basic solution?
 - a. NaC₂H₃O₂
 - NaNO₃
 - c. NH_fCI
 - d. BaSO₄

8. The weakest of the bases listed is

| Acid | Conjugate Base | Ka (Ionization Constant of Acid) | | | | |
|---|--------------------|-------------------------------------|--|--|--|--|
| HCL | CIT | 100% ionized | | | | |
| HSO ₄ T | SO ₄ 2- | 1.2 x 10° | | | | |
| H ₂ S HS ⁻ HS ⁻ S ^{r-} | | 5.7x10 ⁻⁸ | | | | |
| | | 1.2x10 ⁻⁰ | | | | |

- a. CIT
- # 1.00 T
- e. HS⁻ d. S³⁻
- e. 50e²⁻
- 9. The oxide of which element will react with water to form the strongest acid?

| Period | Main Groups | | | | | | | |
|--------|-------------|---|---|----|---|----|-----|------|
| | 1 | п | П | IV | V | VI | VII | 0 |
| First | | 1 | | | | | | |
| Second | X | Y | Z | P | | 9 | S | U |
| Third | WAY | | | | | R | T | 0.40 |

- e. W
- 6. M
- c. P
- d. R.
- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate KC₂H₂O₂, is less alkaline than a 0.10 M solution of potassium evanide. KCN?
 - a. Hydrocyanic acid is a weaker acid than acetic acid.
 - Hydrocyanic acid is less soluble in water than acetic acid.
 - c. Cyanides are less soluble than acetaces
 - d. Acetic acid is a weaker acid than hydrocyanic acid
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of
 - a. a 0.100 M sodium solution.
 - a 0.0500 M sodium hydroxide solution.
 - c. a 0.0500 M benzoic acid solution.
 - a 0.0500 M sodium benzoate solution.

Answer Key For Semester 2 Chemistry Test

Katrin Zwirglmaier

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