

Strong vs. Weak Acids WebLab PhET

Go to: <https://phet.colorado.edu/en/simulation/acid-base-solutions> Click <Introduction> to begin.

Part 1: Procedure

1. The lab has 2 tools that allow you to test for pH values: A probe , and pH paper . Use each one by dipping it into the solution to be tested. Try all the given types of solutions and fill in the Data Chart with the pH value 0-14.

2. The circuit with a battery and bulb as shown:  is the tool used to test for conduction of a solution. By dipping the wire leads into the solution, the bulb will either remain unlit, be dimly lit, be somewhat bright or very bright. Test each solution and record your observation for the bulb's brightness in the chart below.

Part 1: Data	pH Value from Probe	Color & pH Value from pH Paper	Observations from Circuit Tool Describe the brightness
Water (H_2O) 	7.00	Yellow / orange 7	light
Strong Acid (HA) 	2.00	Red / orange 6	Very bright
Weak Acid (A) 	4.50	orange 6	Bright
Strong Base (MOH) 	12.00	Blue / green 11	Very bright
Weak Base (B) 	9.50	Blue / green 10	Bright

Part 1: Analysis

1. What pH value range is observed: a. for acids? 0 - 6 b. for bases? 8 - 14

2. Why are some solutions better conductors of electricity?

The strong electrolytes often supply ions to the solution. This happens more than the weak electrolytes. It is determined by the high and low pH levels.

Answer Key Pogil Strong Versus Weak Acids

Robert A. Richardson



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