

QUESTION 2 (contd)

(b. iii) The system is an exothermic one, when it is cooled, the equilibrium will adjust to annul the cooling effect by releasing more heat. And this favours the forward reaction.

$$\text{(c) i) } \text{PbCl}_2 \rightarrow \text{Pb}^{2+} + 2\text{Cl}^-$$

$$\text{No of moles of PbCl}_2 = \frac{\text{mass}}{\text{molar mass}} = \frac{92.7}{207 + (35.5 \times 2)}$$

$$= \frac{92.7}{278}$$

$$= 0.33345 \text{ moles}$$

$$\text{No of moles of Pb}^{2+} \text{ ionized} = 0.33345 \text{ moles}$$

$$\text{mass of lead present} = \text{no of moles} \times \text{molar mass}$$

$$= 0.33345 \times 207$$

$$= 69.0 \text{ grams}$$

$$\text{(ii) Quantity of electricity used} = It = I \times 4.8 \times 3600$$

$$= (17280 \text{ I}) \text{ C}$$



2×96500 liberate 207 grams of Pb

17280I will liberate $\frac{17280I \times 207}{2 \times 96500} = 207$

$$I = \frac{2 \times 96500 \times 207}{17280 \times 207} = 11.17 \text{ Amperes}$$

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Table of Contents 2014 Waec Answer Sheet Chemistry Obj And Essay

1. Understanding the eBook 2014 Waec Answer Sheet Chemistry Obj And Essay
 - The Rise of Digital Reading 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Advantages of eBooks Over Traditional Books
2. Identifying 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2014 Waec Answer Sheet Chemistry Obj And Essay
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Personalized Recommendations
 - 2014 Waec Answer Sheet Chemistry Obj And Essay User Reviews and Ratings
 - 2014 Waec Answer Sheet Chemistry Obj And Essay and Bestseller Lists

5. Accessing 2014 Waec Answer Sheet Chemistry Obj And Essay Free and Paid eBooks
 - 2014 Waec Answer Sheet Chemistry Obj And Essay Public Domain eBooks
 - 2014 Waec Answer Sheet Chemistry Obj And Essay eBook Subscription Services
 - 2014 Waec Answer Sheet Chemistry Obj And Essay Budget-Friendly Options
6. Navigating 2014 Waec Answer Sheet Chemistry Obj And Essay eBook Formats
 - ePub, PDF, MOBI, and More
 - 2014 Waec Answer Sheet Chemistry Obj And Essay Compatibility with Devices
 - 2014 Waec Answer Sheet Chemistry Obj And Essay Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Highlighting and Note-Taking 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Interactive Elements 2014 Waec Answer Sheet Chemistry Obj And Essay
8. Staying Engaged with 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2014 Waec Answer Sheet Chemistry Obj And Essay
9. Balancing eBooks and Physical Books 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2014 Waec Answer Sheet Chemistry Obj And Essay
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Setting Reading Goals 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Fact-Checking eBook Content of 2014 Waec Answer Sheet Chemistry Obj And Essay
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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