

CHEMISTRY

2c(ii)

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

where:

 V_1 = Initial Volume = 150 cm^3 T_1 = Initial Temp in kelvin = $21^\circ\text{C} + 273 = 294\text{K}$ T_2 = Final Temp in kelvin = $45^\circ\text{C} + 273 = 318\text{K}$ $V_2 = ?$

$$V_2 = V_1 \times \frac{T_2}{T_1}$$

$$V_2 = 150 \text{ cm}^3 \times \frac{318\text{K}}{294\text{K}}$$

$$V_2 = 150 \text{ cm}^3 \times 1.0816$$

$$V_2 = 162.24 \text{ cm}^3$$

2d.

I Sodium oxide:



II Phosphorus (III) oxide:



III Chlorine (VII) oxide



IV Magnesium oxide



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DJ Losen



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Table of Contents 2014 Waec Chemistry Essay And Obj Answers

1. Understanding the eBook 2014 Waec Chemistry Essay And Obj Answers
 - The Rise of Digital Reading 2014 Waec Chemistry Essay And Obj Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 2014 Waec Chemistry Essay And Obj Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2014 Waec Chemistry Essay And Obj Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2014 Waec Chemistry Essay And Obj Answers
 - Personalized Recommendations

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- 2014 Waec Chemistry Essay And Obj Answers and Bestseller Lists
- 5. Accessing 2014 Waec Chemistry Essay And Obj Answers Free and Paid eBooks
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 - 2014 Waec Chemistry Essay And Obj Answers eBook Subscription Services
 - 2014 Waec Chemistry Essay And Obj Answers Budget-Friendly Options
- 6. Navigating 2014 Waec Chemistry Essay And Obj Answers eBook Formats
 - ePub, PDF, MOBI, and More
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 - 2014 Waec Chemistry Essay And Obj Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of 2014 Waec Chemistry Essay And Obj Answers
 - Highlighting and Note-Taking 2014 Waec Chemistry Essay And Obj Answers
 - Interactive Elements 2014 Waec Chemistry Essay And Obj Answers
- 8. Staying Engaged with 2014 Waec Chemistry Essay And Obj Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2014 Waec Chemistry Essay And Obj Answers
- 9. Balancing eBooks and Physical Books 2014 Waec Chemistry Essay And Obj Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2014 Waec Chemistry Essay And Obj Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 2014 Waec Chemistry Essay And Obj Answers
 - Setting Reading Goals 2014 Waec Chemistry Essay And Obj Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 2014 Waec Chemistry Essay And Obj Answers
 - Fact-Checking eBook Content of 2014 Waec Chemistry Essay And Obj Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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