

CHAPTER 4 REVIEW

Arrangement of Electrons in Atoms

SECTION 2

SHORT ANSWER Answer the following questions in the space provided.

1. _____ How many quantum numbers are used to describe the properties of electrons in atomic orbitals?
 (a) 1 (c) 3
 (b) 2 (d) 4
2. _____ A spherical electron cloud surrounding an atomic nucleus would best represent
 (a) an s-orbital. (c) a combination of two different p-orbitals.
 (b) a p-orbital. (d) a combination of an s and a p-orbital.
3. _____ How many electrons can an energy level of $n = 4$ hold?
 (a) 32 (c) 8
 (b) 24 (d) 6
4. _____ How many electrons can an energy level of $n = 2$ hold?
 (a) 32 (c) 8
 (b) 24 (d) 6
5. _____ Compared with an electron for which $n = 2$, an electron for which $n = 4$ has more
 (a) spin. (c) energy.
 (b) particle nature. (d) wave nature.
6. _____ According to Bohr, which is the point in the figure below where electrons cannot reside?
 (a) point A (c) point C
 (b) point B (d) point D



7. _____ According to the quantum theory, point D in the above figure represents:
 (a) the fixed position of an electron.
 (b) the furthest position from the nucleus that an electron can achieve.
 (c) a position where an electron probably exists.
 (d) a position where an electron cannot exist.

Arrangement Of Electrons In Atoms Chapter 4 Review Answers

Ibrahim Sikder



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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