

1. Simplify. Omit restrictions.

a. $\frac{(-xy)^7}{xy^7}$

$\frac{-x^7y^7}{xy^7} = -x^6$

b. $\frac{8a^2b - 12ab^2 + 4ab}{4ab}$

$2a - 3b + 1$

c. $(ab - 2c)(ab + 2c)$ DOTS

$(ab)^2 - (2c)^2$
 $a^2b^2 - 4c^2$

d. $(5p^3 - 6)^2$ PST

$(5p^3)^2 - 2(5p^3)(6) + (6)^2$

$25p^6 - 60p^3 + 36$

e. $(2x - y)(x^2 - 4xy + 3y^2)$

$2x^3 - 8x^2y + 6xy^2 - x^2y + 4xy^2 - 3y^3$

$2x^3 - 9x^2y + 10xy^2 - 3y^3$

f. $\frac{8x - 20y}{4} - \frac{18x + 45y}{9}$

$(2x - 5y) - (2x + 5y)$

$2x - 5y - 2x - 5y$
 $-10y$

2. Factor completely. Show steps.

a. $a^3x - 9ax^3$

$ax(a^2 - 9x^2)$

$ax(a + 3x)(a - 3x)$

b. $5c^2 - 40c - 100$

$5(c^2 - 8c - 20)$

$5(c - 10)(c + 2)$

c. $a^3 + a^2 - a - 1$

$(a^3 + a^2) + (-a - 1)$

$a^2(a + 1) - (a + 1)$

$(a + 1)(a^2 - 1)$
 $(a + 1)(a + 1)(a - 1)$

d. $n^2(n^2 - 9) + 4(9 - n^2)$

$n^2(n^2 - 9) - 4(n^2 - 9)$

$(n^2 - 9)(n^2 - 4)$

$(n + 3)(n - 3)(n + 2)(n - 2)$

e. $36x^2 - 84xy + 49y^2$

$(6x - 7y)^2$

$\begin{matrix} 2ab \\ 2(6x)(7y) \\ 84xy \end{matrix}$

f. $10t^2 - 15t - 25$

$5(2t^2 - 3t - 5)$

$5(2t - 5)(t + 1)$

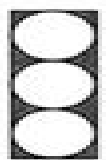
3. If r is the radius of each circle in the diagram at the right, express the area of the shaded region in factored form. Show work.

Rectangle - 3 Circles

$(2r)(6r) - 3\pi r^2$

$12r^2 - 3\pi r^2$

$3r^2(4 - \pi)$



6r

2r

4. Solve and write the solution set. Omit checks.

a. $(n + 3)(2n + 3) = (n + 2)^2 + (n - 2)^2$

$2n^2 + 9n + 9 = (n^2 + 4n + 4) + (n^2 - 4n + 4)$

$2n^2 + 9n + 9 = 2n^2 + 8$

$9n + 9 = 8$

$9n = -1$

$n = -\frac{1}{9}$

$\therefore \{-\frac{1}{9}\}$

b. $(x + 2)(x - 3) = 24$

$x^2 - x - 6 = 24$

$x^2 - x - 30 = 0$

$(x - 6)(x + 5) = 0$

$x = 6$ or $x = -5$

$\therefore \{-5, 6\}$

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N Noddings



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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared Ca(OH)_2 solution with a standard hydrochloric acid solution

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web question experiment 8 determining ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 hci molarity 0.050 M volume of filtered Ca(OH)_2 solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml OH⁻ in saturated Ca(OH)_2 M Ca^{2+} in saturated Ca(OH)_2 M ksp of Ca(OH)_2 average ksp of Ca(OH)_2

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be ksp $\text{Ca}^{2+} \text{OH}^-$ from the

titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for $\text{Ca}(\text{OH})_2$ was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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we calculate the ksp for calcium hydroxide $\text{ksp Ca(OH)}_2 = 5 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error
accepted value for ksp for sodium hydroxide is 5×10^{-6}