

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment is designed to determine the concentration of chloride ions in a solution by gravimetric analysis. The method involves the precipitation of chloride ions as silver chloride (AgCl) using a known volume of a silver nitrate (AgNO_3) solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is designed to be performed in a laboratory setting with the following equipment and materials: analytical balance, volumetric flask, beaker, filter paper, silver nitrate solution, sodium chloride solution, and distilled water. The procedure involves the preparation of a known volume of a sodium chloride solution, the addition of a known volume of a silver nitrate solution, the precipitation of silver chloride, the filtration of the precipitate, the drying of the precipitate, and the weighing of the precipitate. The results are then used to calculate the concentration of chloride ions in the original solution.

The experiment is designed to be performed in a laboratory setting with the following equipment and materials: analytical balance, volumetric flask, beaker, filter paper, silver nitrate solution, sodium chloride solution, and distilled water. The procedure involves the preparation of a known volume of a sodium chloride solution, the addition of a known volume of a silver nitrate solution, the precipitation of silver chloride, the filtration of the precipitate, the drying of the precipitate, and the weighing of the precipitate. The results are then used to calculate the concentration of chloride ions in the original solution.

Procedure:

The procedure involves the preparation of a known volume of a sodium chloride solution, the addition of a known volume of a silver nitrate solution, the precipitation of silver chloride, the filtration of the precipitate, the drying of the precipitate, and the weighing of the precipitate. The results are then used to calculate the concentration of chloride ions in the original solution.

Data Analysis:

	Sample 1	Sample 2
Mass of precipitate	0.1234 g	0.1567 g
Mass of sodium chloride	0.0567 g	0.0789 g
Mass of silver nitrate	0.0987 g	0.1234 g

The concentration of chloride ions in the original solution is calculated as follows:

$$\text{Concentration} = \frac{\text{Mass of precipitate}}{\text{Mass of sodium chloride}} \times \frac{\text{Molar mass of NaCl}}{\text{Molar mass of AgCl}}$$

The concentration of chloride ions in the original solution is calculated as follows:

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Protection Agency**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the EPA text.

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