

9. cobalt (III) ion (Co^{3+}) – 27 electrons as an atom; 24 electrons as an ion.*

[illegible]

2 valence electrons ($4s$) and 1 d electron are removed.

10. chromium (VI) ion (Cr^{6+}) – 24 electrons as an atom; 18 electrons as an ion.

$1s$	$2s$	$2p$	$3s$	$3p$	$4s$	$3d$	$4p$
------	------	------	------	------	------	------	------

2 valence electrons ($4s$) and 4^{th} & 5^{th} electrons are removed

11. arsenide ion (As^{3-}) - 33 electrons as an atom; 36 electrons as an ion.

[illegible]

12. argon (Ar) – 18 electrons

$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$					
1s	2s	2p	3s	3p	4s		3d		4p

13. Use the electron orbital diagrams for questions #2 through #12 to determine how many **valence electrons** are contained in the elements in (a) through (f).

a. #2 Ne 8 b. #3 P 5 * c. #4 Fe 2 *

d. #6 S²⁻ 8 e. #11 As³⁻ 8 f. #12 Ar 8

An **electron orbital diagram** represents each individual electron of an atom or an ion with an arrow to represent electron spin. An **electron configuration** is not as specific of a model. An **electron configuration** represents the sublevels that contain electrons and the number of electrons in each sublevel is represented with a superscript ($X_{\text{number of electrons}}$).

14. beryllium (Be) - 4 electrons*

1s ~~2~~ 2s ~~2~~ 2p — 3s — 3p — 4s — 3d — 4p — 5s — 4d — 5p —

15. chlorine (Cl) – 17 electrons

1s 2 2s 2 2p 6 3s 2 3p 5 4s — 3d — 4p — 5s — 4d — 5p —

16. tellurium (Te) – 52 electrons*

$$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^4$$

Valence Electrons Core Electrons (4)

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Marcel A. Müller



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