

## Chemical Formula Writing Worksheet Solutions

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write "ZnCl<sub>2</sub>", as shown.

|                 | <i>zinc</i>                                                    | <i>iron (II)</i>                                               | <i>iron (III)</i>                                              | <i>gallium</i>                                                 | <i>silver</i>                                     | <i>lead (IV)</i>                                               |
|-----------------|----------------------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------|---------------------------------------------------|----------------------------------------------------------------|
| <i>chloride</i> | <b>ZnCl<sub>2</sub></b>                                        | <b>FeCl<sub>2</sub></b>                                        | <b>FeCl<sub>3</sub></b>                                        | <b>GaCl<sub>3</sub></b>                                        | <b>AgCl</b>                                       | <b>PbCl<sub>4</sub></b>                                        |
| <i>acetate</i>  | <b>Zn(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> | <b>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> | <b>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></b> | <b>Ga(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></b> | <b>Ag C<sub>2</sub>H<sub>3</sub>O<sub>2</sub></b> | <b>Pb(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>4</sub></b> |
| <i>nitrate</i>  | <b>Zn(NO<sub>3</sub>)<sub>2</sub></b>                          | <b>Fe(NO<sub>3</sub>)<sub>2</sub></b>                          | <b>Fe(NO<sub>3</sub>)<sub>3</sub></b>                          | <b>Ga(NO<sub>3</sub>)<sub>3</sub></b>                          | <b>AgNO<sub>3</sub></b>                           | <b>Pb(NO<sub>3</sub>)<sub>4</sub></b>                          |
| <i>oxide</i>    | <b>ZnO</b>                                                     | <b>FeO</b>                                                     | <b>Fe<sub>2</sub>O<sub>3</sub></b>                             | <b>Ga<sub>2</sub>O<sub>3</sub></b>                             | <b>Ag<sub>2</sub>O</b>                            | <b>PbO<sub>2</sub></b>                                         |
| <i>nitride</i>  | <b>Zn<sub>3</sub>N<sub>2</sub></b>                             | <b>Fe<sub>3</sub>N<sub>2</sub></b>                             | <b>FeN</b>                                                     | <b>GaN</b>                                                     | <b>Ag<sub>3</sub>N</b>                            | <b>Pb<sub>3</sub>N<sub>4</sub></b>                             |
| <i>sulfate</i>  | <b>ZnSO<sub>4</sub></b>                                        | <b>FeSO<sub>4</sub></b>                                        | <b>Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>              | <b>Ga<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>              | <b>Ag<sub>2</sub>SO<sub>4</sub></b>               | <b>Pb(SO<sub>4</sub>)<sub>2</sub></b>                          |

Write the formulas for the following compounds:

- 1) copper (II) chloride **CuCl<sub>2</sub>**
- 2) lithium acetate **LiC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>**
- 3) vanadium (III) selenide **V<sub>2</sub>Se<sub>3</sub>**
- 4) manganese (IV) nitride **Mn<sub>3</sub>N<sub>4</sub>**
- 5) beryllium oxide **BeO**
- 6) sodium sulfate **Na<sub>2</sub>SO<sub>4</sub>**
- 7) aluminum arsenide **AlAs**
- 8) potassium permanganate **KMnO<sub>4</sub>**
- 9) chromium (VI) cyanide **Cr(CN)<sub>6</sub>**
- 10) tin (II) sulfite **SnSO<sub>3</sub>**
- 11) vanadium (V) fluoride **VF<sub>5</sub>**
- 12) ammonium nitrate **NH<sub>4</sub>NO<sub>3</sub>**

# Answer Key For Chemical Formula Writing

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Answer Key For Chemical Formula Writing**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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