

GIZMO's - Activity: Moles, Avogadro, and Molar Mass- Answer key 2023

Name: _____

Date: _____

Activity: Moles, Avogadro, and Molar Mass

Prior Knowledge Questions

1. In the image, note a dozen eggs, a dozen donuts and a dozen roses. How many of each item do you have? 12 eggs, 12 donuts, and 12 roses.
2. Would a dozen of each object have the same mass? No
3. Suppose you have a dozen silver atoms, a dozen oxygen atoms, and a dozen lithium atoms. Even though you have the same number of each, will they all have the same mass? Explain.
They will not all have the same mass because although there may be an equal amount of each one, they don't all weigh the same or take up the same amount of space.



Gizmo Warm-up

On the AVOGADRO CONSTANT tab, place the copper (Cu) atom on the nano-balance on the left, which will show the average atomic mass of copper rather than the mass of a single copper atom.



1. What is the average mass of a copper atom? 63.546 u
The unit "u" refers to **unified atomic mass units**. (Officially, 1 u is one-twelfth the mass of a C-12 atom.)
2. To know how many atoms are in a gram of copper, use the larger balance on the right. Press **Add atoms** to put a scoop of atoms in the weighing dish, and keep adding until the balance registers between 1 and 2 grams. If you don't seem to be making much progress, adjust the exponent using the slider, which will make the scoop size bigger.
How many atoms did you need to add? 9.6×10^{21} atoms of copper

Answer Key For Moles And Mass

Oswaal Editorial Board



Answer Key For Moles And Mass:

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isotopes are atoms that have the same number of protons but different numbers of neutrons thus atoms of 6 12 c and of 6 13

c are isotopes of each other

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web what is the percent composition of cu in terms of these two isotopes answer 69 15 cu 63 and 30 85 cu 65 figure

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web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical

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web verified by toppr correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope s mass number

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