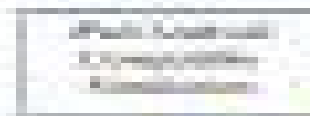


Experiment 1: The Effect of Temperature on the Rate of Reaction

Name: _____

Background Information

Every science experiment is designed to test a hypothesis. In this experiment, the hypothesis is that the rate of reaction will increase as the temperature increases. The rate of reaction is the change in concentration of a reactant or product per unit time. The rate of reaction can be measured by the change in concentration of a reactant or product over a given time interval. The rate of reaction can also be measured by the change in volume of a gas produced over a given time interval. The rate of reaction can also be measured by the change in mass of a solid product over a given time interval.



Procedure

1. Weigh out 1.0 g of sodium hydroxide (NaOH) and place it in a beaker.

2. Add 10 cm³ of water to the beaker.

3. Add 10 cm³ of 2.0 M hydrochloric acid (HCl) to the beaker.

Results

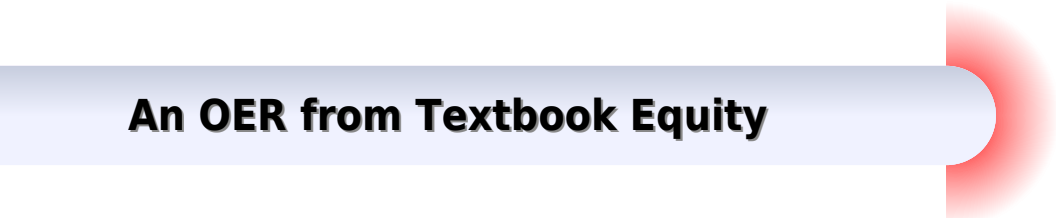
Table 1.1

1. Record the volume of gas produced at regular intervals. (Timing starts when the reaction begins.)
2. Plot a graph of volume of gas produced against time.
3. Calculate the rate of reaction at various intervals.
4. Repeat the experiment at different temperatures. (At least 5 different temperatures.)

Time (s)	Volume of gas (cm ³)	Rate of reaction (cm ³ /s)	Temperature (°C)	Rate of reaction (cm ³ /s)
0	0	0	20	0
10	10	1.0	25	1.2
20	20	1.0	30	1.5
30	30	1.0	35	1.8
40	40	1.0	40	2.0
50	50	1.0	45	2.2
60	60	1.0	50	2.5
70	70	1.0	55	2.8
80	80	1.0	60	3.0
90	90	1.0	65	3.2
100	100	1.0	70	3.5

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