

Name \_\_\_\_\_

Period \_\_\_\_\_

Date Tuesday, November 1, 2016

## Periodic Trends

### ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.
  - a) Al B
  - b) S O
  - c) Br Cl
  - d) Na Al
  - e) O F
  - f) Mg Ca

6. Put the following elements in order from smallest to largest atomic radius **and** explain why:  
C, O, Sn, Sr.

### ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.
  - a) Ca Ga
  - b) Li O
  - c) Cl S
  - d) Br As
  - e) Ba Sr
  - f) O S

### GENERAL QUESTIONS

14. Which group tends to form +1 ions? \_\_\_\_\_
15. Which group tends to form +2 ions? \_\_\_\_\_
16. Which group tends to form -1 ions? \_\_\_\_\_
17. Which group tends not to form ions or react? \_\_\_\_\_

# Answer Key To Periodic Trends

**The Princeton Review**



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## **Table of Contents Answer Key To Periodic Trends**

1. Understanding the eBook Answer Key To Periodic Trends
  - The Rise of Digital Reading Answer Key To Periodic Trends
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Periodic Trends
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Key To Periodic Trends
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Periodic Trends
  - Personalized Recommendations
  - Answer Key To Periodic Trends User Reviews and Ratings

- Answer Key To Periodic Trends and Bestseller Lists
- 5. Accessing Answer Key To Periodic Trends Free and Paid eBooks
  - Answer Key To Periodic Trends Public Domain eBooks
  - Answer Key To Periodic Trends eBook Subscription Services
  - Answer Key To Periodic Trends Budget-Friendly Options
- 6. Navigating Answer Key To Periodic Trends eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer Key To Periodic Trends Compatibility with Devices
  - Answer Key To Periodic Trends Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Key To Periodic Trends
  - Highlighting and Note-Taking Answer Key To Periodic Trends
  - Interactive Elements Answer Key To Periodic Trends
- 8. Staying Engaged with Answer Key To Periodic Trends
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Key To Periodic Trends
- 9. Balancing eBooks and Physical Books Answer Key To Periodic Trends
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Key To Periodic Trends
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key To Periodic Trends
  - Setting Reading Goals Answer Key To Periodic Trends
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key To Periodic Trends
  - Fact-Checking eBook Content of Answer Key To Periodic Trends
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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