

- 27. It is necessary to know the formula of a compound in order to calculate its percentage composition.
- 28. If the percent by mass of carbon in methane, CH_4 , is 75%, then 100 grams of methane contains 75 g of carbon and 25 grams of hydrogen.
- 29. The formula for methane, CH_4 , is both a molecular and an empirical formula.
- 30. The empirical formula of glucose, $C_6H_{12}O_6$, is CH_2O .

C. Problems

Solve the following problems on the space provided. Show your work.

- 31. How many moles of H_2 in 5.1 g. of $20^{16}O$ atoms of H_2 ?
- 32. How many representative particles are in 2.33 g. $20^{16}O$ and $5H_2$?
- 33. What is the gram molecular mass of methane, CH_4 ?
- 34. Find the mass of each of the following:
 - a. 2.40×10^{23} and 4.00×10^{23}
 - b. 2.40 and 4.00
 - c. 4.00×10^{23} and 2.40
 - d. 4.00 and 2.40

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