#### Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- D CH CH
- 6) C<sub>2</sub>H<sub>11</sub> C<sub>2</sub>H<sub>2</sub>
- 7) WO. WO.
- 8) C2H4O2 CH4O
- 9) X<sub>21</sub>Y<sub>12</sub> X<sub>3</sub>Y
- 6) A compound with an empirical formula of C<sub>2</sub>OH<sub>4</sub> and a molar mass of 88 grams per mole. What is the molecular formula of this compound? C<sub>4</sub>O<sub>2</sub>H<sub>4</sub>.
- 7) A compound with an empirical formula of C<sub>s</sub>H<sub>s</sub>O and a molar mass of 136 grams per mole. What is the molecular formula of this compound? C<sub>s</sub>H<sub>s</sub>O<sub>s</sub>
- B) A compound with an empirical formula of CFBrO and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound? C,F,Br,O,
- 9) A compound with an empirical formula of C<sub>2</sub>H<sub>6</sub>N and a molar mass of 46 grams per mole. What is the molecular formula of this compound? C<sub>2</sub>H<sub>6</sub>N
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula C<sub>2</sub>H<sub>2</sub>NO. If its molar mass is 116.1 g/mol, what is the molecular formula of the compound? C<sub>2</sub>H<sub>2</sub>N<sub>3</sub>O<sub>3</sub>.
- 12. A certain blue solid contains 36.84% N and 63.16% ○. What is the empirical formula of this compound? The ratios are N<sub>1.00</sub>O<sub>1.50</sub>. Since 1.50 is not close to a whole number, we multiply both subscripts by 2. The empirical formula is thus N<sub>2</sub>O<sub>3</sub>. (The name is dinitrogen trioxide.)
- A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound? InCl.
- 14. An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula? K<sub>2</sub>C<sub>2</sub>O<sub>4</sub>.
- Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of subbing alcohol? C<sub>2</sub>H<sub>4</sub>O

# **Answer Key For Chemical Formulas**

Y Pai

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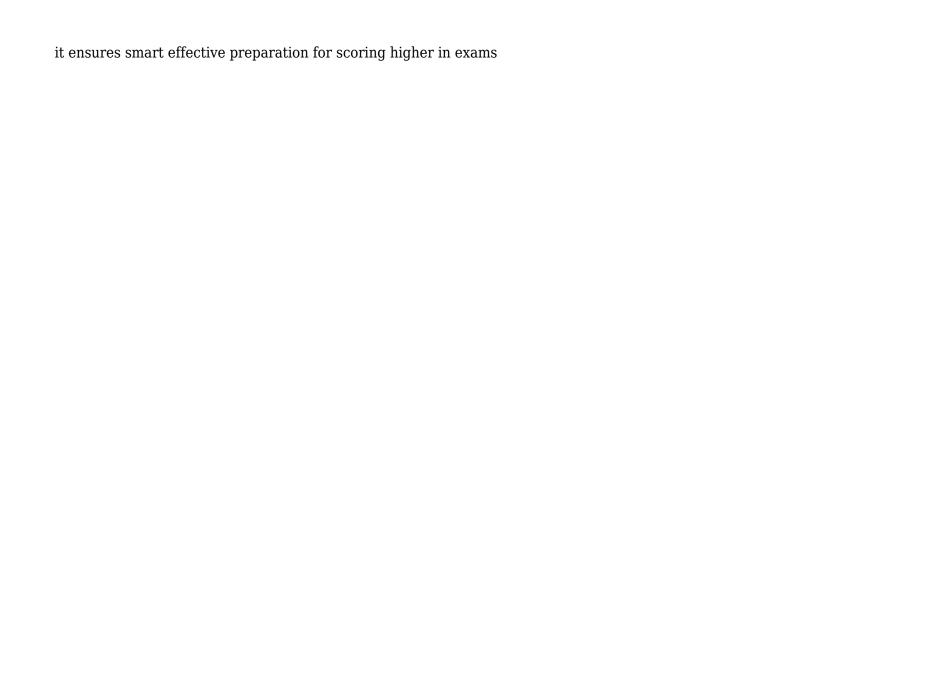
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