

Balancing redox equations

Example

Balance the equation showing the oxidation of Fe^{2+} ions to Fe^{3+} ions by dichromate ions ($\text{Cr}_2\text{O}_7^{2-}$) in an acidic medium



Step 1: Identify oxidising and reducing agents and write half reactions



$$2\text{Cr} = 12$$

$$\text{Cr} = +6$$



Iron loses one electron, i.e. it is *oxidised*. It donates one electron to dichromate, i.e. it is the *reducing agent*.

Half equation: Oxidation reaction: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$



Chromium gains three electrons, i.e. it is *reduced*. It gains its electrons from iron, i.e. it is the *oxidising agent*.

Half equation: Reduction reaction: $\text{Cr}_2^{6+} + 3\text{e}^- \rightarrow \text{Cr}^{3+}$

Analyzing Oxidation Reduction Reactions Answers

M Carnoy



Analyzing Oxidation Reduction Reactions Answers:

Analyzing Oxidation Reduction Reactions Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Analyzing Oxidation Reduction Reactions Answers**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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