

14. You are given a bottle of saturated calcium chloride solution at 25°C. Outline the *experimental procedure* you would use to determine the *solubility* of CaCl_2 in grams/L. (3 marks)

- weigh empty dish
 - add (10.0 mL) of sat. CaCl_2 soln.
 - evaporate
 - weigh solid CaCl_2 (s)
 - Calculate sol. in $\frac{g}{L} = g \div L$

15. A solution of FeBr_3 has a $[\text{Br}^-]$ which is 0.36 M. Find the $[\text{Fe}^{3+}]$ in the same solution. (1 mark)

Answer 0.12 M

16. You dissolve 3.20×10^{-2} moles of $\text{Fe}(\text{NO}_3)_3$ into water to make 2.00 L of solution.

a) Determine the $[\text{Fe}^{3+}]$ (1 mark)

Answer 0.0160 M ($1.60 \times 10^{-2} \text{ M}$)

b) Determine the $[\text{NO}_3^-]$ (1 mark)

Answer 0.0480 M ($4.80 \times 10^{-2} \text{ M}$)

17. 30.0 mL of 0.850 M $\text{Ba}(\text{NO}_3)_2$ is added to 50.0 mL of water. Pay attention to S.D.s in this question!

a) Find the new $[\text{Ba}(\text{NO}_3)_2]$ in the solution. (2 marks)

$$0.850 \text{ M} \times \frac{30.0 \text{ mL}}{80.0 \text{ mL}} = 0.31875 \text{ M}$$

Answer 0.319 M

b) Find the new $[\text{NO}_3^-]$ in the solution. (1 mark)

Answer 0.638 M

18. 67.0 mL of 0.25 M BaCl_2 is mixed with 25.0 mL of 0.30 M NaCl . Determine the final $[\text{Cl}^-]$ in the new mixture. (2 marks)

$$[\text{BaCl}_2] = 0.25 \text{ M} \times \frac{67.0 \text{ mL}}{92.0 \text{ mL}} = 0.182065 \text{ M} \quad \begin{matrix} [\text{Cl}^-] = 0.36413 \text{ M} \\ [\text{Cl}^-] = 0.08152 \text{ M} \end{matrix}$$

$$[\text{NaCl}] = 0.30 \text{ M} \times \frac{25.0}{92.0} = 0.08152 \text{ M}$$

Answer 0.45 M

Answer Key For Mastering Chemistry 1

Barbara Herlihy



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