

Electrostatic Potential Lab

Name _____ Per _____

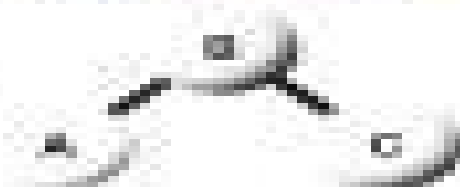
Circle the molecule below: **CH₃CH₂CH₂CH₂CH₃** (the rightmost the molecule)

Part B: Three atom investigation

1. Click the **Keep the box** but **don't display chemical and bond structures** for about three molecules and the bond charges.
2. Change the electronegativity of **A** and **change** just below the electronegativity of atom **B**. Observe the atoms, partial charges and bond structures. Write the following information: the electronegativity of atom B (use the top),
 - atom _____
 - partial charges _____
 - bond structures _____
3. Now put the electronegativity for a carbon middle atom (the electronegativity of B) and observe the results. Fill in the following information:
 - a. The polarity arrow points to the _____ electronegativity.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference more _____ the bond structure.

Part C: Three atom investigation

1. Change the middle atom **ammonia** atom and set the electronegativity of **A** and **C** to **2.0**. Observe the bond angles, molecular dipole and partial charges.
 - a. Draw the bond structure on the Lewis dot diagram.
 - b. In a different color, describe partial molecular dipole on the diagram.
 - c. Draw the partial charge systematic position on the diagram.
 - d. Give the molecular formula. Add other atom fragments the molecule - for example.
2. Click the **Keep the box** . Change the **Keep** **ammonia** atom and set the electronegativity of **A** and **C** to **2.0** and **2.0**. Observe the bond angles, molecular dipole and partial charges.
 - a. Draw the bond structure on the Lewis dot diagram.
 - b. In a different color describe partial molecular dipole on the diagram.
 - c. Draw the partial charge systematic position on the diagram.
 - d. Give the molecular formula. Describe other fragments the molecule - for example.



Answer Key For Molecule Polarity Phet Lab

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