

Practice Sheet

Electron Dot (Lewis) Structures

A Lewis or Electron Dot Structure is a convenient representation of the valence electrons in an atom.

An electron dot structure for an atom is simply the symbol for the element, surrounded by a number of dots equal to the number of valence electrons.

Avoid a common mistake: the dots represent valence electrons only, so make sure you use only the number of dots corresponding to the number of valence electrons.

Examples:



- Draw Lewis dot structures for an atom of each of following elements:

1. K

3. Ar

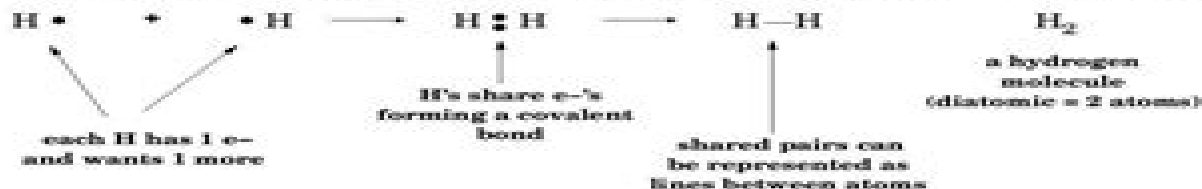
2. Si

4. As

Covalent Bonds

An atom can also achieve a noble gas configuration by sharing electrons with other atoms. When two non-metals combine, they typically share electrons in covalent bonds and form what are known as covalent compounds. We can draw Lewis Electron Dot Diagrams for covalent molecules to determine what type of bonds the molecules form.

For example, consider two hydrogen atoms. Each one has one valence electron. However, both hydrogens would like to have two electrons, to obtain the same configuration as [He]. To do this, each H atom shares an electron with the other, forming a covalent bond. In H_2 , each H atom has 2 electrons.



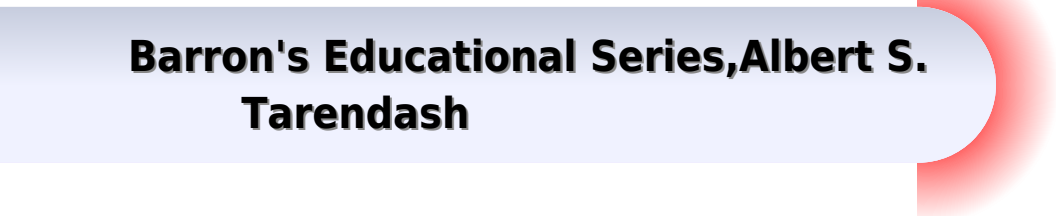
Here's another example, this time for chlorine. A Cl atom has 7 valence electrons, so it needs 1 more to achieve an octet (8 electrons).



Each chlorine now has 8 electrons, because you can count the shared ones:

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