

ChemQuest 39

Concentration

Name: _____

Date: _____

Hour: _____

Information: Molarity

Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

Critical Thinking Questions

- Consider the terms "concentrated" and "dilute". Are these qualitative or quantitative terms?
These are qualitative terms; very general and do not include the magnitude or quantity.
- One way of quantitatively measuring solution concentration is with units of molarity, symbolized by M. You see 1.7 liters (L) of a sodium chloride and water solution. The label on the bottle reads "1.5 M NaCl". You don't know what molarity is, but you decide to find out. After evaporating the water out of the solution you discover that there are about 149 grams of salt. Using this information, which of the following formulas is/are correct for finding molarity?

$$A) \text{ Molarity} = \frac{\text{grams of solute}}{\text{moles of solute}}$$

$$149\text{g} \div 58.5\text{g/mol} = 2.547 \text{ mol}$$

$$B) \text{ Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

$$1.5 = 2.547 \text{ mol} \div 1.7 \text{ L}$$

- Using the equation you discovered in question two, calculate the molarity of each of the following solutions.
A) A solution is prepared by dissolving 24.9 g of CaCl_2 in 210 mL (which 0.210 L) of solution.

$$1.07 \text{ M}$$

$$\text{Change g to mol: } 24.9\text{g} \div 111.1\text{g/mol} = 0.224 \text{ mol}$$

$$\text{Molarity} = \text{mol} \div \text{L} = 0.224 \div 0.210 = 1.07 \text{ M}$$

- A solution contains 12.9 g of Na_2SO_4 in 325 mL of solution.

$$0.279 \text{ M}$$

$$\text{Change g to mol: } 12.9 \div 142.1\text{g/mol} = 0.09078 \text{ mol}$$

$$\text{Molarity} = \text{mol} \div \text{L} = 0.09078\text{mol} \div 0.325\text{L} = 0.279 \text{ M}$$

- Verify that I need 2.15 moles of $\text{Ca}(\text{NO}_3)_2$ to make 358 mL of a 6.00 molar solution.

$$\text{mol} = (\text{Molarity})(\text{Liters}) = (6.00\text{M})(0.358\text{L}) = 2.15 \text{ mol}$$

- Verify that it takes 80.8 g of sodium chloride to make 425 mL of a 3.25 M solution.

$$\text{mol} = (\text{Molarity})(\text{Liters}) = (3.25\text{M})(0.425\text{L}) = 1.38 \text{ mol}$$

$$(1.38 \text{ mol})(58.5\text{g/mol}) = 80.7 \text{ g}$$

Answer Key To Chemquest 39 Concentration

IM Harris



Answer Key To Chemquest 39 Concentration:

Answer Key To Chemquest 39 Concentration: Bestsellers in 2023 The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous engrossing novels captivating the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the engaging narratives that have captivated audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has gripped readers with its raw and emotional exploration of domestic abuse. Hoover expertly weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can prevail. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This intriguing historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Answer Key To Chemquest 39 Concentration : Delia Owens "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting. These top-selling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and gripping novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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