



Colligative Properties

Answer Key

Vocabulary: boiling point, colligative property, concentration, dissociate, freezing point, manometer, osmosis, osmotic pressure, solute, solution, solvent, vapor pressure

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

1. When the forecast calls for freezing rain, salt is often spread on the roads. Why is this done?

Salt is spread on the roads to prevent ice and snow from accumulating.

2. Antifreeze is added to car radiators to prevent the water in the radiators from both freezing and boiling. How do you think antifreeze affects the temperatures at which water freezes (the **freezing point**) and boils (the **boiling point**)?

Antifreeze lowers the freezing point and raises the boiling point of water.

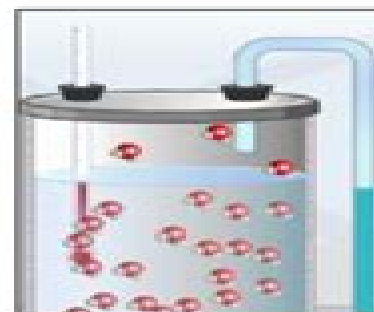
Gizmo Warm-up

Adding salt or other substances to water can affect the temperature at which it freezes or boils. These effects and others, known as **colligative properties**, are explored in the Colligative Properties Gizmo.

Check that **No solute** is selected and the **Air temp.** is 25 °C.

1. Describe the motion of the water molecules in the beaker.

The water molecules move randomly in the container. When water molecules collide, they bounce off of one another like rubber balls.



2. Change the **Air temp.** to -10 °C. What do you see as the water freezes?

As molecules freeze from the surface down, they stop moving freely and begin to vibrate in place.

3. Now change the **Air temp.** to 110 °C. What do you see as the water boils?

The molecules move more rapidly and bubbles appear in the beaker. The water disappears from the U-shaped tube that is attached to the beaker.



Answer Key 18 3 Colligative

Oswaal Editorial Board

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Physical Chemistry for Engineering and Applied Sciences Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto Designed to be as rigorous as compatible with a first year student's ability to understand the text presents detailed step by step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed In addition to the 600 exercises and end of chapter problems the text is rich in worked non trivial examples many of which are designed to be inspiring and thought provoking Step by step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes The answers to all of the 112 worked examples 99 exercises following many of the worked examples and 496 end of chapter problems Topics not normally seen in introductory physical chemistry textbooks ionic reaction rates activities and activity coefficients or not regularly explained in much detail electrochemistry chemical kinetics with an eye on industrial applications Special appendices that provide detailed explanations of basic integration and natural logarithms for students lacking a background in integral calculus An in depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations

Study Guide for Chemical Principles Thomas Elliott Taylor, 1979

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