

Example Problem 10-3**Converting Mass to Moles**

A roll of copper wire has a mass of 545 g. How many moles of copper are in the roll? Use the atomic mass of copper given on the periodic table to apply a conversion factor to the mass given in the problem.

$$\text{moles of Cu} = \text{grams Cu} \times \frac{1 \text{ mol Cu}}{\text{grams Cu}}$$

$$\text{moles of Cu} = 545 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.546 \text{ g Cu}} = 8.58 \text{ mol Cu}$$

Note that the conversion factor is arranged so that g Cu cancels, leaving only mol Cu, the desired quantity.

Example Problem 10-4**Converting Moles to Mass**

Calculate the mass of 0.025 moles of calcium.

Use the molar mass of calcium to apply a conversion factor to the number of moles given in the problem. According to the periodic table, the atomic mass of calcium is 40.078 amu, so the molar mass of calcium is 40.078 g.

$$\text{mass of Ca} = \text{mol Ca} \times \frac{\text{grams Ca}}{1 \text{ mol Ca}}$$

$$\text{mass of Ca} = 0.025 \text{ mol Ca} \times \frac{40.078 \text{ g Ca}}{1 \text{ mol Ca}} = 1.00 \text{ g Ca}$$

The conversion factor is arranged so that mol Ca cancels, leaving g Ca, the desired quantity.

Practice Problems

- Calculate the mass of 0.88 mol selenium.
- A chemist needs 0.0780 mol selenium for a reaction. What mass of selenium should the chemist use?
- A sample of sulfur has a mass of 225 g. How many moles of sulfur are in the sample?

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