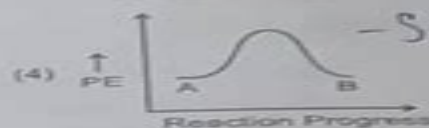


SECTION - C

Previous Years Questions

1. For a reaction $A \rightarrow B$, enthalpy of reaction is -4.2 kJ mol^{-1} and enthalpy of activation is 9.6 kJ mol^{-1} . The correct potential energy profile for the reaction is shown in option

[NEET-2021]



$$\Delta H = -4.2 \text{ kJ}$$

$$\Delta H = E_{\text{for}} - E_{\text{back}}$$

$$-4.2 = 9.6 - E_{\text{back}}$$

$$E_{\text{back}} = 9.6 + 4.2$$

$$E_{\text{back}} = 13.8$$

$$-5 \times 10^3 = \frac{E_a}{2.303RT}$$

$$-5 \times 10^3 = \frac{E_a}{2.303 \times 8.314 \times R}$$

$$E_a = 5000 \times 2.303 \times 8.314 \times R$$

$$E_a = 1151618.314$$

2. The slope of Arrhenius plot ($\ln k$ vs $\frac{1}{T}$) of first order reaction is $-5 \times 10^3 \text{ K}$. The value of E_a of the reaction is. Choose the correct option for your answer.

[Given $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$]

[NEET-2021]

- (1) -83 kJ mol^{-1} (2) 41.5 kJ mol^{-1}
(3) 83.0 kJ mol^{-1} (4) 166 kJ mol^{-1}

3. The rate constant for a first order reaction is $4.606 \times 10^{-3} \text{ s}^{-1}$. The time required to reduce 2.0 g of the reactant to 0.2 g is

[NEET-2020 (Phase-1)]

- (1) 200 s (2) 500 s
(3) 1000 s (4) 100 s

4. An increase in the concentration of the reactants of a reaction leads to change in

[NEET-2020 (Phase-1)]

- (1) Heat of reaction
(2) Threshold energy
(3) Collision frequency
(4) Activation energy

5. The half-life for a zero order reaction is 0.02 M initial concentration of reactant is. The rate constant (in $\text{mol L}^{-1} \text{ s}^{-1}$) for the reaction is

[NEET-2020 (Phase-1)]

- (1) 1.0×10^{-2}
(2) 1.0×10^{-4}
(3) 2.0×10^{-4}
(4) 2.0×10^{-5}

$$9.6$$

$$4.2$$

$$13.8$$

In collision theory of chemical reaction represents

[NEET-2020 (Phase-1)]

- (1) The fraction of molecules with energy greater than E_a
(2) The fraction of molecules with energy greater than E_a
(3) The collision frequency of reactants
(4) Steric factor

A first order reaction has a rate constant $2.303 \times 10^{-3} \text{ s}^{-1}$. The time required for the reactant to reduce to 10 g will be [Given that $\log_2 = 0.3010$]

[NEET-2020 (Phase-1)]

- (1) 602 s (2) 230.3 s
(3) 301 s (4) 2000 s

8. For a reaction, activation energy is $1.6 \times 10^5 \text{ J mol}^{-1}$. The rate constant at 400 K will be

[Given that gas constant, $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$]

[NEET-2020 (Phase-1)]

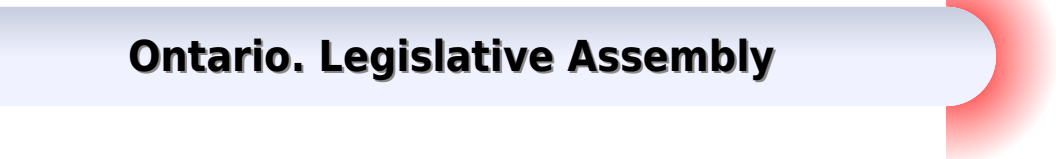
- (1) $3.2 \times 10^5 \text{ s}^{-1}$ (2) $3.2 \times 10^6 \text{ s}^{-1}$
(3) $1.6 \times 10^5 \text{ s}^{-1}$ (4) $1.6 \times 10^6 \text{ s}^{-1}$

9. If the rate constant for a first order reaction is k , the time (t) required for 99% of the reaction is given by

- (1) $t = 0.693/k$ (2) $t = 0.693/k$
(3) $t = 4.606/k$ (4) $t = 4.606/k$

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