

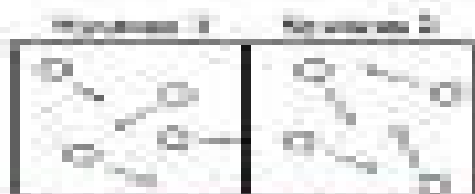
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with m in cm^2/s^2 is the mass of particles and molecules mass and change, directly connected as all of temperature increase. The Boltzmann distribution curve shows that the average velocity of all of the distribution's molecules (area under the curve) is proportional to temperature.

3. Sketch a curve to represent the relative velocities of the particles in the following two gases, samples of the same gas.



- a. In System 1, is the average kinetic temperature more System 1 or System 2? Justify your answer.

System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out than in System 2, so the particles in System 1 have more energy. According to the equation $KE = \frac{1}{2}mv^2$, the particles in System 1 have a higher average velocity than the particles in System 2, so the particles in System 1 are moving faster than the particles in System 2.

- b. Are either attracted with kinetic energy for interaction? Justify your answer in terms of molecular collisions.

Yes, both are attracted to each other. In System 1, the particles are more spread out than in System 2, so the particles in System 2 are more spread out than the particles in System 1. This means that the particles in System 2 are more spread out than the particles in System 1, so the particles in System 2 are more spread out than the particles in System 1. This means that the particles in System 2 are more spread out than the particles in System 1, so the particles in System 2 are more spread out than the particles in System 1.

Answer Key Unit 6 Thermochemistry

Sherman F. Booth



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