



## General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

### Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{P_1}{P_2} = \frac{Z_1}{Z_2}$$

$$P_{\text{solution}} = \chi_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = \chi_{\text{solute}} P_{\text{solvent}}^{\circ}$$

### Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \quad (\text{at } 25^{\circ}\text{C})$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592\text{V}}{n} \log(K) \quad (\text{at } 25^{\circ}\text{C})$$

**USE ONE:**

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} - E_{\text{anode(re.d.)}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} + E_{\text{anode(ox.)}}^{\circ}$$

### Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surr}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

$$\Delta S_{\text{ran}}^{\circ} = \Sigma S_{\text{products}}^{\circ} + \Sigma S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{ran}}^{\circ} = \Sigma \Delta G_{\text{products}}^{\circ} + \Sigma \Delta G_{\text{reactants}}^{\circ}$$

### Concentrations

$$\text{molality} = \frac{\text{moles solute}}{\text{mass solvent (kg)}}$$

$$\% \text{ by mass} = \frac{\text{moles solute}}{\text{moles solute} + \text{moles solvent}} \times 100\%$$

$$\text{parts per million} = \frac{\text{moles solute}}{\text{moles solute} + \text{moles solvent}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{moles solute}}{\text{moles solute} + \text{moles solvent}} \times 10^9$$

### Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_a + pK_b = 14.00 \quad (\text{at } 25^{\circ}\text{C})$$

$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$$pOH = pK_b + \log \frac{[BH^+]}{[B]}$$

$$\% \text{ ionization} = \frac{[H_3O^+]_{\text{eq}}}{[HA]_{\text{eq}}} \times 100\%$$

$$pK_a = -\log(K_a)$$

$$pK_b = -\log(K_b)$$

$$K_a = 10^{-pK_a}$$

$$K_b = 10^{-pK_b}$$

### Equilibria

$$Q_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[A]^x[B]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k} \quad (\text{1st order reactions only})$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

# Answer Sheet General Chemistry Equations

**Stanley Wallace Morse**



## **Answer Sheet General Chemistry Equations:**

## Whispering the Techniques of Language: An Mental Journey through **Answer Sheet General Chemistry Equations**

In a digitally-driven earth wherever screens reign supreme and immediate communication drowns out the subtleties of language, the profound secrets and mental nuances concealed within words often go unheard. However, situated within the pages of **Answer Sheet General Chemistry Equations** a fascinating fictional prize sporting with fresh thoughts, lies an exceptional quest waiting to be undertaken. Published by a talented wordsmith, that wonderful opus encourages visitors on an introspective trip, gently unraveling the veiled truths and profound influence resonating within the very cloth of each and every word. Within the psychological depths with this moving evaluation, we shall embark upon a heartfelt exploration of the book is primary subjects, dissect their captivating writing fashion, and fail to the strong resonance it evokes heavy within the recesses of readers hearts.

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