

SCHAU Unit 4 - Quiz – Equilibrium

Player(s): _____

Period(s): _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Useful formulae:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad ; \quad K_p = \frac{p_{\text{products}}}{p_{\text{reactants}}}$$

$$K_c = \frac{p_{\text{products}}}{p_{\text{reactants}}} \quad ; \quad K_p = \frac{p_{\text{products}}}{p_{\text{reactants}}} \cdot \frac{p_{\text{total}}}{p_{\text{total}}}$$

gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$; $p_{\text{total}} = p_{\text{gas1}} + p_{\text{gas2}} + \dots + p_{\text{gasn}}$; $n_{\text{gas}} = n_{\text{gas1}} + n_{\text{gas2}} + \dots + n_{\text{gasn}}$

$K_p = K_c \cdot (RT)^{\Delta n}$

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000 K, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.020 atm. The reaction is:



If the volume of an 8.00 L flask at 1000 K where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of H_2S at equilibrium?
- What is the value of K_c at equilibrium?
- What other equilibria have chemical formula involving H_2 and S_2 at equilibrium above 1000 K?

	$\text{H}_2(\text{g})$	$\text{S}_2(\text{g})$	$\text{H}_2\text{S}(\text{g})$
Initial concentration (mole/L)			
Change in concentration (mole/L)			
Equilibrium concentration (mole/L)	0.020	0.010	0.020

What is the value of K_c at equilibrium?

Answers For Sch4u Unit 4

Tsfx



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Table of Contents Answers For Sch4u Unit 4

1. Understanding the eBook Answers For Sch4u Unit 4
 - The Rise of Digital Reading Answers For Sch4u Unit 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Sch4u Unit 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Sch4u Unit 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Sch4u Unit 4
 - Personalized Recommendations
 - Answers For Sch4u Unit 4 User Reviews and Ratings
 - Answers For Sch4u Unit 4 and Bestseller Lists
5. Accessing Answers For Sch4u Unit 4 Free and Paid eBooks
 - Answers For Sch4u Unit 4 Public Domain eBooks
 - Answers For Sch4u Unit 4 eBook Subscription Services
 - Answers For Sch4u Unit 4 Budget-Friendly Options
6. Navigating Answers For Sch4u Unit 4 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Sch4u Unit 4 Compatibility with Devices
 - Answers For Sch4u Unit 4 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Sch4u Unit 4
 - Highlighting and Note-Taking Answers For Sch4u Unit 4
 - Interactive Elements Answers For Sch4u Unit 4
8. Staying Engaged with Answers For Sch4u Unit 4

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answers For Sch4u Unit 4
- 9. Balancing eBooks and Physical Books Answers For Sch4u Unit 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Sch4u Unit 4
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Sch4u Unit 4
 - Setting Reading Goals Answers For Sch4u Unit 4
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Sch4u Unit 4
 - Fact-Checking eBook Content of Answers For Sch4u Unit 4
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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