

9 ChemActivity 41-Alternate Solubility and the Solubility Product

6. $\text{PbCl}_2(\text{s})$ is not very soluble in water.



- If x moles of $\text{PbCl}_2(\text{s})$ dissolve in 1.00 L of water, how many moles of $\text{Pb}^{2+}(\text{aq})$ are produced? How many moles of $\text{Cl}^{-}(\text{aq})$ are produced?
- The equilibrium constant, K_{sp} , for the dissolution of $\text{PbCl}_2(\text{s})$ in water is 1.6×10^{-5} . What is the concentration of $\text{Pb}^{2+}(\text{aq})$ at equilibrium? What is the concentration of $\text{Cl}^{-}(\text{aq})$ at equilibrium?

7. For each of the following situations, determine whether or not a precipitate of MgF_2 is expected to form.

- 500.0 mL of 0.050M $\text{Mg}(\text{NO}_3)_2$ is mixed with 500.0 mL of 0.010 M NaF.
- 500.0 mL of 0.050M $\text{Mg}(\text{NO}_3)_2$ is mixed with 500.0 mL of 0.0010 M NaF.

Answers To Chemactivity 41

**Michael P. Garoutte, Ashley B.
Mahoney**



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Organic Chemistry, a Guided Inquiry Andrei Straumanis, 2004 *Chemistry* Richard S. Moog, John J. Farrell, 2017-06-26

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