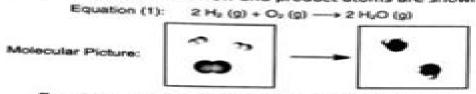
# Balancing Chemical Equations

(What stays the same and what may change in a chemical reaction?)

# Model 1: Atoms are conserved in chemical reactions

Atoms are neither created nor destroyed when chemical reactions occur. Two balanced chemical reactions (or equations) are given below. In a chemical reaction equation, reactant atoms are shown on the left side of the arrow and product atoms are shown on the right.



# Equation (2): $Fe_1O_3$ (s) + 2 Al (s) $\longrightarrow$ 2 Fe (l) + Al<sub>2</sub>O<sub>3</sub> (s)

# Critical Thinking Questions:

- Considering Equation (1) above:
  - a. What are the reactants? Write your answer in both words and formulas. Check that your teammates agree.
  - b. What is the product? Write your team's answer in both words and formulas.
- 2. What does the arrow represent in a chemical reaction?
- 3. How many hydrogen molecules are represented on the reactant side of equation (1)?
- 4. How many water molecules are represented on the product side of equation (1)?
- 5. How many H atoms are represented on the reactant side of equation (1)? On the product side? \_\_\_\_
- 6. How many O atoms are represented on the reactant side of equation (1)? On the product side? \_
- 7. Compare the number of atoms of each element (H and O) on the reactant and product sides of equation (1). Are these numbers the same or different (circle one)?
- 8. Compare the number of atoms of each element (Fe, Al, and O) on the reactant and product sides of equation (2). Are these numbers the same or different (circle one37
- In reaction (1):
  - a. For each H<sub>2</sub> molecule consumed, how many H<sub>2</sub>O molecules are produced?
  - b. For each O2 molecule consumed, how many H2O molecules are produced?
  - c. Recall that a mole is simply a specific large quantity, like a dozen or a gross. For each mole of oxygen molecules consumed, how many moles of water molecules are produced? \_\_\_\_

# **Answers To Chemactivity 45**

**Michael P. Garoutte** 

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#### **Table of Contents Answers To Chemactivity 45**

- 1. Understanding the eBook Answers To Chemactivity 45
  - The Rise of Digital Reading Answers To Chemactivity 45
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Answers To Chemactivity 45
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers To Chemactivity 45
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answers To Chemactivity 45
  - Personalized Recommendations

- Answers To Chemactivity 45 User Reviews and Ratings
- Answers To Chemactivity 45 and Bestseller Lists
- 5. Accessing Answers To Chemactivity 45 Free and Paid eBooks
  - Answers To Chemactivity 45 Public Domain eBooks
  - Answers To Chemactivity 45 eBook Subscription Services
  - Answers To Chemactivity 45 Budget-Friendly Options
- 6. Navigating Answers To Chemactivity 45 eBook Formats
  - o ePub, PDF, MOBI, and More
  - Answers To Chemactivity 45 Compatibility with Devices
  - Answers To Chemactivity 45 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers To Chemactivity 45
  - Highlighting and Note-Taking Answers To Chemactivity 45
  - Interactive Elements Answers To Chemactivity 45
- 8. Staying Engaged with Answers To Chemactivity 45
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers To Chemactivity 45
- 9. Balancing eBooks and Physical Books Answers To Chemactivity 45
  - $\circ$  Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers To Chemactivity 45
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answers To Chemactivity 45
  - Setting Reading Goals Answers To Chemactivity 45
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers To Chemactivity 45
  - Fact-Checking eBook Content of Answers To Chemactivity 45

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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