

1 What is 6050.287 rounded to the nearest ten?

- A 6050
- B 6100
- C 6050.29
- D 6050.3

CS00001204

2 The total land area for the United States is 3,537,438 square miles. What is this value rounded to the nearest thousand square miles?

- A 3,500,000
- B 3,537,000
- C 3,538,000
- D 3,540,000

CS00001071

3 What is 40% of 250?

- A 50
- B 100
- C 150
- D 200

CS00001070

4 What is $\frac{3}{8}$ written as a percent?

- A 26.7%
- B 30%
- C 37.5%
- D 50%

CS00001012

5 What is the decimal 0.7 written as a fraction?

- A $\frac{1}{7}$
- B $\frac{3}{4}$
- C $\frac{3}{7}$
- D $\frac{7}{10}$

CS00000007

6 A company donated 200 books to a local library. If 70 of them are fiction, what percent of the donated books are fiction?

- A 35%
- B 40%
- C 60%
- D 65%

CS00000040

7 In a parking lot, 1 out of every 8 cars is blue. What percent of the cars in this lot are blue?

- A 1.25%
- B 7%
- C 9%
- D 12.5%

CS00000000

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Harcourt Brace



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

$$8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$$

 149.4 g H_2O is your theoretical yield so percent yield = $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2/\text{mole}} = 0.666 \text{ moles } \text{O}_2$$

$$\frac{21.3 \text{ g } \text{CO}}{28 \text{ g } \text{CO}/\text{mole}} = 0.761 \text{ moles } \text{CO}$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

 b determine the theoretical yield of Al_2O_3

$$\frac{160.0 \text{ g } \text{Al}}{27 \text{ g } \text{Al}/\text{mole}} = 5.93 \text{ moles } \text{Al}$$

$$\frac{5.93 \text{ moles } \text{Al}}{4 \text{ moles } \text{Al}} \times 2 \text{ moles } \text{Al}_2\text{O}_3 \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mole} = 302.3 \text{ g } \text{Al}_2\text{O}_3$$

 c determine the percent yield

$$\frac{260.0 \text{ g } \text{Al}_2\text{O}_3}{302.3 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 86.01\%$$

 theoretical = 302.3 g

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