

Zumdahl review questions Ans

40. 2.07 g MgO

41.

42. To determine the limiting reactant, first calculate the number of moles of each reactant present. Then determine how these number of moles corresponds to the mole ratio indicated by the balanced equation.

43.

44. A reactant is present in excess if there is more of it present than is needed to combine with the limiting reactant for the process. An excess doesn't affect the theoretical yield.

45.

46. a. Limiting HCl; 18.3 g AlCl_3 ; 0.415 g H_2
b. Limiting NaOH; 19.9 g Na_2CO_3 ; 3.38 g H_2O
c. Limiting $\text{Pb}(\text{NO}_3)_2$; 12.6 g Pb; 5.71 g HNO_3
d. Limiting I_2 ; 19.6 g KI

47

48. a. Limiting Na; 84.9 g NaNH_2
b. Limiting BaCl_2 ; 56.0 g BaSO_4
c. Limiting NaOH; 78.8 g Na_2SO_3
d. Limiting H_2SO_4 ; 58.1 g $\text{Al}_2(\text{SO}_4)_3$

49.

50. a. Limiting CO_2 ; 11.4 mg CH_3OF
b. Limiting I_2 ; 10.7 mg AlI_3
c. Limiting HBr; 12.4 mg CaBr_2 ; 2.23 mg H_2O
d. Limiting H_3PO_4 ; 15.0 mg CrPO_4 ; 0.309 mg H_2

51.

52. 136 g urea

53.

54. 14.5 g FeCl_3 produced; 0.5 g Cl_2 remains. Fe limits

55.

56. 0.626 CuI; 0.690 g KI; 0.537 g K_2SO_4

57.

58. 0.67 kg SiC

59.

60. If the reaction occurs in a solvent, the product may have a substantial solubility in the solvent the reaction may come to equilibrium before the full yield of product is achieved; loss of product may occur through operator error.

61.

62. 94.60% yield

63

83.

84. 56.3 g Br_2

85.

86. 0.0771 g H_2

87

88. a. Limiting Br_2 ; 6.4 g NaBr
b. Limiting CuSO_4 ; 5.1 g ZnSO_4 ; 2.0 g Cu
c. Limiting NH_4Cl ; 1.6 g NH_3 ; 1.7 g H_2O ; 5.5 g NaCl
d. Limiting Fe_2O_3 ; 3.5 g Fe; 4.1 g CO_2

89.

90. 0.624 mol N_2 ; 17.5 g N_2 ; 1.25 mol H_2O ; 22.5 g H_2O

91.

92. 5 g

Answers To Review Questions Zumdahl

S Nieto



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