

IDEAL GAS LAW

Name: _____

Use the Ideal Gas Law below to solve the following problems:

$$PV = nRT$$

where P is pressure in atmospheres
 V is volume in liters
 n is number of moles of gas
 R is Ideal Gas Constant
 $0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$
 T is Kelvin temperature

pressure = 1 atmosphere
 volume = 1.0 liter

1. How many moles of oxygen will occupy a volume of 2.5 liters at 1.0 atm and 25°C?
0.104 mol $(2.5 \text{ L}) / (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 0.104 \text{ mol}$

2. What volume will 0.50 mole of nitrogen occupy at 2.00 atm and 25°C?
5.06 L $(0.50 \text{ mol}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 5.06 \text{ L}$

3. What pressure will be exerted by 25.0 g of CO_2 at a temperature of 25°C held in a volume of 0.500 L?
2.82 atm 2 moles $(25.0 \text{ g}) / (44.01 \text{ g/mol}) = 0.568 \text{ mol}$
2.82 atm $(0.568 \text{ mol}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 2.82 \text{ atm}$

4. At what temperature will 0.500 g of O_2 exert a pressure of 1.00 atm held in a volume of 1.00 L?
12.2 K $(0.0156 \text{ mol}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 1.00 \text{ atm}$
12.2 K $(0.0156 \text{ mol}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 1.00 \text{ atm}$

5. What is the density of air, which has an O_2 and N_2 at 1.0 atm and 25°C?
1.19 g/L $(1.0 \text{ atm}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 0.0821 \text{ L/mol}$ $(0.0821 \text{ L/mol}) \cdot (298 \text{ K}) = 24.4 \text{ L/mol}$
1.19 g/L $(29 \text{ g/mol}) / (24.4 \text{ L/mol}) = 1.19 \text{ g/L}$

6. If the density of nitrogen is 1.16 g/L at 1.0 atm and 25°C, what is its molar mass?
28.0 g/mol $(1.16 \text{ g/L}) \cdot (24.4 \text{ L/mol}) = 28.3 \text{ g/mol}$

7. How many moles of nitrogen will occupy a volume of 0.50 L at 2.00 atm and 25°C?
0.050 mol $(0.50 \text{ L}) \cdot (2.00 \text{ atm}) / (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 0.050 \text{ mol}$

8. What volume will 0.500 g of O_2 at a temperature of 1.00 atm and 25°C?
22.4 L $(0.0156 \text{ mol}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 1.00 \text{ atm}$ $(0.0156 \text{ mol}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 1.00 \text{ atm}$

9. Find the number of grams of CO_2 that exert a pressure of 1.00 atm in a volume of 0.50 L at a temperature of 25°C.
22.4 g $(0.50 \text{ L}) \cdot (1.00 \text{ atm}) / (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 0.0156 \text{ mol}$ $(0.0156 \text{ mol}) \cdot (44.01 \text{ g/mol}) = 0.686 \text{ g}$

10. A gas exerts a pressure of 1.00 atm in a volume of 0.50 L at a temperature of 25°C. What is the molar mass of the gas?
28.0 g/mol $(1.00 \text{ atm}) \cdot (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) = 0.0821 \text{ L/mol}$ $(0.0821 \text{ L/mol}) \cdot (298 \text{ K}) = 24.4 \text{ L/mol}$
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Teacher's Name: _____

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