Science

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(Chapter - 1) (Chemical Reactions and Equations) (Intext)

(Class 10)

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Question 1:

A solution of a substance 'X' is used for white washing.

(i) Name the substance 'X' and write its formula.

(ii) Write the reaction of the substance 'X' named in (i) above with water.

Answer 1:

(i) The substance 'X' is calcium oxide. Its chemical formula is CaO.

(ii) Calcium oxide reacts vigorously with water to form calcium hydroxide (slaked lime)

$$CaO + H_2O \longrightarrow Ca(OH)_2$$
Calcium Oxide

Quick Lime

Calcium Hydroxide

Sladed Lime

Question 2:

Why is the amount of gas collected in one of the test tubes in Activity 1.7 double of the amount collected in the other? Name this gas.

_Answer 2:

During the Electrolysis of water, hydrogen and oxygen is get separated by the electricity. Water (H₂O) contains two parts hydrogen and one part oxygen. Since hydrogen goes to one test tube and oxygen goes to another, the amount of gas collected in one of the test tubes is double of the amount collected in the other.

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Question 1:

Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

_Answer 1:

When an iron nail is dipped in a copper sulphate solution, iron (which is more reactive than copper) displaces copper from copper sulphate solution forming iron sulphate, which is green in colour.

Therefore, the blue colour of copper sulphate solution fades and green colour appears.

Question 2:

Give an example of a double displacement reaction other than the one given in Activity 1.10.

_Answer 2:

Sodium carbonate reacts with calcium chloride to form calcium carbonate and sodium chloride.

In this reaction, sodium carbonate and calcium chloride exchange ions to form two new compounds. Hence, it is a double displacement reaction.

Question 3:

Identify the substances that are oxidised and the substances that are reduced in the following reactions.

(iii)
$$CuO(s) + H_2(g) \rightarrow Cu(s) + H_2O(l)$$

(i) Sodium (Na) is oxidised as it gains oxygen and oxygen gets reduced.

(ii) Copper oxide (CuO) is reduced to copper (Cu) while hydrogen (H₂) gets oxidised to water (H₂O).

Answers To Mastering Chemistry Ch 1 Questions

Pasquale De Marco

Answers To Mastering Chemistry Ch 1 Questions:

CBSE Class 12 Mastering Organic Reactions MCQ With Answers Priti Singhal, 2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students educators and researchers by focusing on reaction mechanisms practical applications and problem solving techniques Organic chemistry is not just about memorizing equations and formulas it is about understanding how molecules interact change and influence each other under different conditions With that in mind this book emphasizes the logic and patterns behind organic reactions making it easier for readers to apply concepts across a variety of scenarios Each chapter of this book builds upon foundational knowledge ensuring a progressive learning experience From nucleophilic substitutions to pericyclic reactions and from oxidation reduction mechanisms to named reactions we cover both fundamental and advanced topics to cater to students at all levels Real world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals biochemistry agriculture and environmental science This approach bridges the gap between theory and practical applications helping readers appreciate the relevance of organic chemistry in daily life Chemistry in Action: Strategies and Techniques for Mastering the Advanced Placement Examination Pasquale De Marco, 2025-07-20 Chemistry in Action Strategies and Techniques for Mastering the Advanced Placement Examination is the ultimate resource for students preparing for the Advanced Placement Chemistry examination With its comprehensive coverage of topics engaging explanations and abundant practice opportunities this book sets students on the path to success Inside Chemistry in Action Strategies and Techniques for Mastering the Advanced Placement Examination readers will find a wealth of valuable content tailored to the AP Chemistry curriculum Clear and concise explanations of fundamental concepts are complemented by numerous examples illustrations and practice problems These elements work together to foster a deep understanding of chemical principles and prepare students for the rigors of the AP Chemistry exam Beyond its focus on exam preparation Chemistry in Action Strategies and Techniques for Mastering the Advanced Placement Examination also provides a broader perspective on the field of chemistry It delves into the practical applications of chemistry in various industries showcasing how chemical principles impact our everyday lives This approach not only enhances students appreciation for the subject but also equips them with a foundation for future studies or careers in chemistry related fields More than just a textbook Chemistry in Action Strategies and Techniques for Mastering the Advanced Placement Examination is a journey of discovery Through engaging storytelling and thought provoking questions it invites students to explore the wonders of chemistry and unravel the mysteries of the natural world This book transforms the study of chemistry into an immersive and enjoyable experience inspiring a lifelong passion for scientific exploration With its comprehensive content engaging writing style and abundance of practice opportunities Chemistry in Action Strategies and Techniques for Mastering the Advanced Placement Examination is the perfect companion for students seeking to excel in the AP Chemistry exam and beyond It is a valuable resource for high

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