REVIEW OF BASIC CHEMISTRY ANSWER KEY

Name the following elements. Spelling counts:

Write the symbols for the following elements.

H	hydrogen
S	sulphur
CI	chlorine
C	carbon
N	nitrogen
Na	sodium
Pb	lead
Ag	silver

sodium	PMan
phosphorus	P
fluorine	-
magnesium	Mg
potassium	K
calcium	Car
zina	Zn
iron	Fo

Using a periodic table, record the atomic number for the following elements:

gold (Au) 79 copper (Cu) 29

 Using a periodic table, record the atomic mass for the following elements to one decimal place:

chlorine (CI) 35.5 calcium (Ca) 40.1

 Complete the following table. Use the information provided in the chart – not a periodic table – to determine atomic masses, but you may refer to a periodic table to name the element.

Element	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons
Lead	82	207	82	125	82
Argon	18	58	18	40	19
Barium	56	137	56	81	56
Tin	50	112	50	62	50

6. Match the unit with its description:

_F	302	amount of substance	A.	joule
C	2.	density	в.	g
В	3.	mass	G.	g/mL
D	4.	molar mass	D.	g/mok
A_	5.	energy	€.	L
E	6.	volume	F.	mole

Answers To 2nd First Semester Chemistry Review

Disha Experts

Answers To 2nd First Semester Chemistry Review:

Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books The teacher is one who guides students through the subject matter helps each student stay on schedule and be organized and is their source of accountability along the way With that in mind this guide provides additional help through the laboratory exercises as well as lessons guizzes and examinations that are provided along with the answers The lessons in this study emphasize working through procedures and problem solving by learning patterns The vocabulary is kept at the essential level Practice exercises are given with their answers so that the patterns can be used in problem solving These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college Guided labs are provided to enhance instruction of weekly lessons There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions It is important to see the hand of God and His principles and wisdom as it plays out in chemistry This course integrates what God has told us in the context of this study Features Each suggested weekly schedule has five easy to manage lessons that combine reading and worksheets Worksheets guizzes and tests are perforated and three hole punched materials are easy to tear out hand out grade and store Adjust the schedule and materials needed to best work within your educational program Space is given for assignments dates. There is flexibility in scheduling Adapt the days to your school schedule Workflow Students will read the pages in their book and then complete each section of the teacher guide They should be encouraged to complete as many of the activities and projects as possible as well Tests are given at regular intervals with space to record each grade About the Author DR DENNIS ENGLIN earned his bachelor's from Westmont College his master of science from California State University and his EdD from the University of Southern California He enjoys teaching animal biology vertebrate biology wildlife biology organismic biology and astronomy at The Master's University His professional memberships include the Creation Research Society the American Fisheries Association Southern California Academy of Sciences Yellowstone Association and Au Sable Institute of Environmental Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a Studies carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Organic Chemistry, Student Study Guide and Solutions Manual David

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