

6. Consider 670 mL of a 4.10 M solution of $\text{Mg}(\text{NO}_3)_2$ setting in a beaker. If you evaporate all 670 mL of the solution, how many grams of solute would be left in the beaker?

407.8g

$\text{mol} = (\text{Molarity})(\text{Liters}) = (4.10\text{M})(0.670\text{L}) = 2.747 \text{ mol}$
 convert mol to g: $(2.747\text{mol})(148.3\text{g/mol}) = \mathbf{407.8\text{g}}$

Information: Molality

Molality is another way of expressing solution concentration. The symbol for molality is m.

Whereas molarity (M) represents the ratio of moles solute to liters of solution, the molality (m) is the ratio of moles solute to kilograms of solvent. It can be expressed using the following formula:

$$\text{molality} = \frac{\text{moles of solute}}{\text{kg solvent}}$$

Critical Thinking Questions

7. Consider a solution that is prepared by adding 1.34 moles of sodium nitrate to 2.5 kg of water. What is the molality of the solution?

0.536 m

$m = 1.34\text{mol} \div 2.5\text{kg} = 0.536 \text{ m}$

8. Considering the data given in question 7, is this enough data to find the molarity? If so, calculate the molarity. If not, explain why not.

No, because we don't know the total volume of the solution.

9. What is the molality of a solution that is made by dissolving 32.6 g of Na_2SO_4 in 475 g of water?

0.483 m

Convert g to mol: $32.6\text{g} \div 142.1\text{g/mol} = 0.229 \text{ mol}$
 $\text{molality} = \text{mol/kg} = 0.229\text{mol} \div 0.475\text{kg} = \mathbf{0.483 \text{ m}}$

10. Consider 2.35 moles of sodium chloride are dissolved in 1.21 kg of solution to make 1.29 liters. Calculate and compare the molarity and molality.

Molarity: 1.82 M

$M = \text{mol} \div L = 2.35\text{mol} \div 1.29 \text{ L} = 1.82 \text{ M}$

molality: 1.94 m

$m = \text{mol} \div \text{kg} = 2.35\text{mol} \div 1.21 \text{ kg} = 1.94 \text{ m}$

11. If 26.45g of Na_2SO_4 are dissolved in 1.10 kg of solution to make 1.24 L, calculate both the molarity and the molality of the resulting solution.

Convert g to mol: $26.45\text{g} \div 142.1\text{g/mol} = 0.186\text{mol}$

Molarity: 0.150 M

$M = \text{mol} \div L = 0.186\text{mol} \div 1.24 \text{ L} = 0.150 \text{ M}$

molality: 0.169 m

$m = \text{mol} \div \text{kg} = 0.186\text{mol} \div 1.10\text{kg} = 0.169 \text{ m}$

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