

35. What is the minimum wavelength of a photon of light that can excite an electron in the hydrogen atom from the $n = 1$ to the $n = 8$ energy level?

$$\Delta E = -2.178 \times 10^{-18} \text{ J} \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} \left(\frac{1}{8^2} - \frac{1}{1^2} \right)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} \left(\frac{1}{64} - 1 \right)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} (0.015625 - 1)$$

$$\Delta E = 2.14 \times 10^{-18} \text{ J}$$

$$\Delta E = \frac{hc}{\lambda} \quad \text{therefore} \quad \lambda = \frac{hc}{\Delta E}$$

$$\lambda = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s})(2.9979 \times 10^8 \text{ m/s})}{(2.14 \times 10^{-18} \text{ J})}$$

$$\lambda = 9.28 \times 10^{-8} \text{ m} = 92.8 \text{ nm}$$

36. In the hydrogen spectrum, what is the wavelength of light associated with the $n = 2$ to $n = 1$ electron transition?

$$\Delta E = -2.178 \times 10^{-18} \text{ J} (1/n_f^2 - 1/n_i^2)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} (1/1^2 - 1/2^2)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} (1/1 - 1/4)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} (1 - 0.25)$$

$$\Delta E = -2.178 \times 10^{-18} \text{ J} (0.75)$$

$$\Delta E = -1.63 \times 10^{-18} \text{ J}$$

$$E = h\nu \text{ and } c = \lambda\nu \quad \text{so } E = hc/\lambda \quad \text{and} \quad \lambda = hc/E$$

$$\lambda = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s})(2.9979 \times 10^8 \text{ m/s})}{1.63 \times 10^{-18} \text{ J}}$$

$$\lambda = 1.22 \times 10^{-7} \text{ m}$$

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