

Chemistry Worksheet: Gas Stoichiometry

For full credit you must show all work and box your final answer.

Write a balanced chemical equation when necessary.

1. What mass of urea is produced from 9.0 liters of ammonia at a temperature of 23°C and a pressure of 14.0 atm? The reaction is $2\text{NH}_{3(g)} + \text{CO}_{2(g)} \rightarrow \text{H}_2\text{NCONH}_{2(l)} + \text{H}_2\text{O}_{(l)}$

$$n_{\text{NH}_3} = 5.19 \text{ mol} \cdot \frac{1 \text{ mol H}_2\text{NCONH}_2}{2 \text{ mol NH}_3} \cdot \frac{60 \text{ g}}{1 \text{ mol}} = \frac{156 \text{ g}}{160 \text{ g}}$$

2. What volume of oxygen gas is needed to completely react with 14.2 g of graphite at a temperature of 740°C and at a pressure of 550 kPa? The reaction is $\text{C}_{(s)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)}$

$$\downarrow \begin{matrix} 5.43 \\ \text{atm} \end{matrix} \quad 14.2 \text{ g} \cdot \frac{1 \text{ mol}}{12 \text{ g}} \cdot \frac{1 \text{ mol O}_2}{1 \text{ mol C}} = n_{\text{O}_2} = 1.18 \text{ mol} \xrightarrow{PV=nRT} 18 \text{ L} = V$$

3. Zinc metal reacts with hydrochloric acid to produce hydrogen gas and zinc chloride. What mass of zinc is needed to produce 50.0 ml of hydrogen gas at 21°C and 112 kPa?



4. As we have seen in the lab, potassium metal reacts explosively with water by producing hydrogen gas and leaving potassium hydroxide in solution. What volume of hydrogen gas is produced when a 0.50 g piece of potassium is added to water at 26.5°C and 1.014 atm pressure?



5. In the combustion of propane gas, C_3H_8 , what volume of carbon dioxide gas is produced when a 20.0 lbs. (9.1 kg) tank is completely consumed at 26.0°C and 1.0 atm of pressure?



6. When strongly heated, CaCO_3 decomposes into carbon dioxide gas and calcium oxide. What mass of calcium carbonate is needed to produce 800.0 ml of carbon dioxide at 19.0°C and 0.966 atm?



Answers To Gas Stoichiometry Worksheet

Sebastian Brünink

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