

Study Card: The Periodic Law

Mendeleev's Periodic Law

In 1869, Russian chemist Dmitri Mendeleev wrote the names of the 63 known elements on cards and ordered them according to mass in one long row. He observed:

When elements are arranged in order of increasing atomic mass, one observes a periodic repetition of their chemical and physical properties.



Mendeleev's Periodic Table

Struck by this periodic repetition in properties, Mendeleev organized elements with identical properties into columns (a.k.a., groups). The result was one of the first Periodic Tables. While Mendeleev's table had only seven groups (there were only 63 elements), many of the elements in those groups still persist in the modern Periodic Table.

Group 1 Group 2 Group 13 Group 14 Group 15 Group 16 Group 17

Atomic Number ... Not Mass

A few elements troubled Mendeleev, most notably, Te and I. When ordered by mass, these two elements seemed out of place. Mendeleev was certain the mass of Te had been measured incorrectly. Since 1869, we have discovered the atomic number. Elements are now ordered by atomic number.

52
Te
Tellurium
127.60

53
I
Iodine
126.90

The Periodic Law

Today the Periodic Law is stated as:

When elements are arranged in order of increasing atomic number, one observes a periodic repetition of their chemical and physical properties.

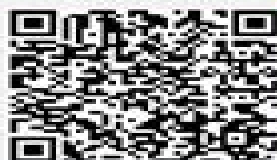
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Modern Periodic Table

(<https://pubchem.ncbi.nlm.nih.gov/periodic-table/>)

Answers To Review The Periodic Law

E Staff



Answers To Review The Periodic Law:

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1. This book is structured into several chapters, namely:
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 - Chapter 2: Essential Elements of Answers To Review The Periodic Law
 - Chapter 3: Answers To Review The Periodic Law in Everyday Life
 - Chapter 4: Answers To Review The Periodic Law in Specific Contexts
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