

Answers to Selected Exercises

The answers listed here are from the *Complete Solutions Guide*, in which rounding is carried out at each intermediate step in a calculation in order to show the correct number of significant figures for that step. Therefore, an answer given here may differ in the last digit from the result obtained by carrying extra digits throughout the entire calculation and rounding at the end (the procedure you should follow).

Chapter 1

19. A law summarizes what happens, e.g., law of conservation of mass, is a chemical reaction or the ideal gas law, $PV = nRT$. A theory (model) is an attempt to explain why something happens. Dalton's atomic theory explains why mass is conserved in a chemical reaction. The kinetic molecular theory explains why pressure and volume are inversely related at constant temperature and moles of gas. 21. The fundamental steps are: (1) making observations; (2) formulating hypotheses; (3) performing experiments to test the hypotheses. The key to the scientific method is performing experiments to test hypotheses. If after the test of time, the hypotheses seem to account satisfactorily for some aspect of natural behavior, then the set of tested hypotheses turns into a theory (model). However, scientists continue to perform experiments to refine or replace existing theories. 23. A qualitative observation expresses what makes something what it is; it does not involve a number, e.g., the air we breathe is a mixture of gases, for it less dense than water, rotten milk stinks. The SI units are mass in grams, length in meters, and volume in the derived units of m^3 . The assumed uncertainty in a number is ± 1 in the last significant figure of the number. The precision of an instrument is related to the number of significant figures associated with an experimental reading on that instrument. Different instruments for measuring mass, length, or volume have varying degrees of precision. Some instruments only give a few significant figures for a measurement while others will give more significant figures. 25. Significant figures are the digits we associate with a number. They contain all of the certain digits and the first uncertain digit (the first estimated digit). What follows is one thousand indicated to varying numbers of significant figures: 1000 or 1×10^3 (1 S.F.); 1.0×10^3 (2 S.F.); 1.00×10^3 (3 S.F.); 1000. or 1.000×10^3 (4 S.F.). To perform the calculation, the addition/subtraction significant rule is applied to $5.5 - 1.0$. The result of this is the one significant figure answer of 0.5. Next, the multiplication/division rule is applied to $0.5/0.50$. A one significant number divided by a two significant number yields an answer with one significant figure (answer = 1). 27. The slope of the T_p vs. T_b plot is 1.8 ($= 9/5$) and the y-intercept is 32°F . The slope of T_p vs. T_b plot is 1 and the y-intercept is -32°C . 29. The density of a gas is much smaller than the density of a solid or a liquid. The molecules in a solid and a liquid are very close together, whereas in the gas phase, the molecules are very far apart. 31. a. exact; b. inexact; c. exact; d. inexact. 33. a. 3; b. 4; c. 4; d. 5; e. 3; f. 1; g. 3; h. 3. 35. a. 3.42×10^{-4} ; b. 1.004×10^3 ; c. 1.7992×10^2 ; d. 3.37×10^2 . 37. $2.85 + 0.280 = 3.13$ mL; the graduated cylinder on the left limits the precision of the total volume; it is the least precise measuring device between the two graduated cylinders. 39. a. 640.0; b. 1.327; c. 77.34; d. 3215; e. 0.420. 41. a. 188.3; b. 12; c. 4×10^{-7} ; d. 6.3×10^{-15} ; e. 4.9; Uncertainty appears in the first decimal place. The average of several numbers can be only as precise as the least precise number. Averages can be exceptions to the significant figure rules. f. 0.22; 43. a. 84.3 mm; b. 2.40 m; c. 2.945×10^{-3} cm; d. 1.445 km ; e. 2.393×10^3 mm; $0.00035 \mu\text{m}$. 45. a. 8 lb and 9.9 oz; $20\frac{1}{2}$ lb; b. 4.0×10^3 m; 4.0×10^3 m; c. 1.2×10^{-11} m³; 12 L; 750 lb_f; 0.42 lb_f. 47. a. 4.00×10^3 rods; 10.0 furlongs; 2.01×10^3 m; 2.01 km; b. 8390.0 rods; 209.25 furlongs; 42,195 m; 42,195 km. 49. a. 0.333 kg; 0.822 lb; b. 31.1 g; 136 grams; c. 19.3 cm^3 . 55. 24 capsules. 53. 2.91×10^6 knots; $3.36 \times$

10^6 min. 55. 100 ft. 57. 1.8 g. 59. 7×10^3 kg mercury. 61. a. -273°C , 0 K; b. -40°C , 233 K; c. 20°C , 293 K; d. $4 \times 10^2^\circ\text{C}$, 4×10^2 K. 63. a. 312.4 K; 102.2°F ; b. 248 K; -15°F ; c. 0 K; -459°F ; d. 1024 K; 1470°F . 65. $168^\circ\text{C} = 320^\circ\text{F}$; 67°C ; $T_c = (T_F + 10^\circ\text{C})\frac{5}{9}$; b. 11.4°C ; c. 120°C ; 425 K; 309°F . 69. b. with float density = 0.80 g/cm^3 . 71. $1 \times 10^3 \text{ g/cm}^3$. 73. a. 0.28 cm^3 ; b. 49 carats. 75. 3.8 g/cm^3 . 77. a. Both are the same mass; b. 1.0 mL mercury; c. Both are the same mass; d. 1.0 L benzene. 79. a. 1.0 kg feather; b. 100 g water; c. same. 81. 2.77 cm. 83. a. Picture iv represents a gaseous compound. Pictures ii and iii also contain a gaseous compound but have a gaseous element present. b. Picture vi represents a mixture of two gaseous elements. c. Picture v represents a solid element. d. Pictures ii and iii both represent a mixture of a gaseous element and a gaseous compound. 85. a. heterogeneous; b. homogeneous (hopefully); c. homogeneous; d. heterogeneous (hopefully); e. heterogeneous; f. heterogeneous. 87. a. pure; b. mixture; c. mixture; d. pure; e. mixture (copper and zinc); f. pure; g. mixture; h. mixture; i. mixture. Iron and uranium are elements. Water is a compound. Table salt is usually a homogeneous mixture composed mostly of sodium chloride (NaCl), but it usually will contain other substances that help absorb water vapor (an anticaking agent). 89. Compound. 91. a. physical; b. chemical; c. physical; d. chemical. 93. The sphere that floats has the smaller density, this will be the object that has the same mass in a larger volume. So the object with the larger diameter will have the larger volume and will have the smaller density; the larger sphere will float. 95. 0.010 kg ; 97.10×10^3 bags. 99. 3.0×10^7 m. 101. a. 1.1 g; b. 4.9 g; 103. 14 g. 105. 50.50°C . 107. a. Volume $\times$ density = mass; the orange block is more dense. Since mass (orange) $>$ mass (blue) and volume (orange) $<$ volume (blue), the density of the orange block must be greater to account for the larger mass of the orange block. b. Which block is more dense cannot be determined. Since mass (orange) $>$ mass (blue) and volume (orange) $>$ volume (blue), the density of the orange block may or may not be larger than the blue block. If the blue block is more dense, then its density cannot be so large that the mass of the smaller blue block becomes larger than the orange block. mass. c. The blue block is more dense. Since mass (blue) = mass (orange) and volume (blue) $<$ volume (orange), the density of the blue block must be larger to equate the masses. d. The blue block is more dense. Since mass (blue) $>$ mass (orange) and the volumes are equal, the density of the blue block must be larger to give the blue block the larger mass. 109. $8.5 \pm 0.5 \text{ g/cm}^3$. 111. 3.548×10^3 cable lengths; 5.671×10^3 m; 3048 nautical miles. 113. Phosphorus would be a liquid. 115. a. False; sugar is generally considered to be the pure compound sucrose, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$. b. False; elements and compounds are pure substances. c. True; air is a mixture of mostly nitrogen and oxygen gases. d. False; gasoline has many additives, so it is a mixture. e. True; compounds are broken down to elements by chemical change. 117. In a subtraction, the result gets smaller, but the uncertainties add. If the two numbers are very close together, the uncertainty may be larger than the result. For example, let us assume we want to take the difference of the following two measured quantities, 999.999 ± 2 and 999.995 ± 2 . The difference is 4 ± 4 . Because of the larger uncertainty, subtracting two similar numbers is bad practice. 119. a. 3%; b. 2.2%; c. 0.2%. 121. $d_{\text{max}} = 8.8 \text{ g/cm}^3$, $d_{\text{min}} = 7.17 \text{ g/cm}^3$, $d_{\text{max}}/d_{\text{min}} = \text{max } m_{\text{max}}/\text{min } m_{\text{min}} = 0.81$. The difference in mass is accounted for by the difference in the alloy used (if the assumptions are correct). 123. 7.0%. 125. a. One possibility is that rope B is not attached to anything and rope A and rope C are connected via a pair of pulleys and/or gears. b. Try to pull rope B out of the box. Measure the distance moved by C for a given movement of A. Hold either A or C firmly while pulling on the other rope. 127. $\$160/\text{person}$; 3.20×10^6 nickels/person;

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G Thomas



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A Textbook of Algebraic Number Theory Sudesh Kaur Khanduja, 2022-04-26 This self contained and comprehensive textbook of algebraic number theory is useful for advanced undergraduate and graduate students of mathematics The book discusses proofs of almost all basic significant theorems of algebraic number theory including Dedekind's theorem on splitting of primes Dirichlet's unit theorem Minkowski's convex body theorem Dedekind's discriminant theorem Hermite's theorem on discriminant Dirichlet's class number formula and Dirichlet's theorem on primes in arithmetic progressions A few research problems arising out of these results are mentioned together with the progress made in the direction of each problem Following the classical approach of Dedekind's theory of ideals the book aims at arousing the reader's interest in the current research being held in the subject area It not only proves basic results but pairs them with recent developments making the book relevant and thought provoking Historical notes are given at various places Featured with numerous related exercises and examples this book is of significant value to students and researchers associated with the field The book also is suitable for independent study The only prerequisite is basic knowledge of abstract algebra and elementary number theory

Have You Seen My... Umm... Memory? Miller Caldwell, 2009 Everyone has had the frustrating experience of putting something down and then forgetting where Although there are many factors that contribute to the ability to remember it is also a simple fact that some people are better at remembering than others Outside influences like stress hormone problems or even the side effects of medications can make people absentminded and forgetful But what if your memory is simply getting progressively worse Are you frequently embarrassed at not being able to remember someone's name Have you ever driven away from a petrol station and genuinely forgotten to pay Are these episodes of memory loss driving you to despair If so then this book by Miller Caldwell is not only going to amuse and entertain you but it will provide you with many serious tips and selfhelp exercises that will help you train your brain all over again Mr Caldwell begins by explaining his own diagnosed condition called mild cognitive impairment MCI and the frustration he felt at having to retire early because of shortterm memory loss He explains in simple terms how the human brain processes our thoughts and then stores them in preparation for recall He explains how in most people sharp thinking and reasoning skills are unaffected by shortterm memory decline and suggests that if an older person learns information in the right way he or she is just as likely to remember it as a younger counterpart Each helpful tip and suggested exercise comes with an anecdote from Caldwell's many varied and fascinating life experiences including how he almost married the wrong woman Presented in a delightful and straightforward format this book will be an invaluable companion to those of us who are just a little bit forgetful or equally to those in the early stages of dementia Oh and just in case you forget the charming cover illustration of an elephant with a knot in its trunk will remind you to pick up a copy of your own

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Table of Contents Answers To Selected Exercises Umm

1. Understanding the eBook Answers To Selected Exercises Umm
 - The Rise of Digital Reading Answers To Selected Exercises Umm
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Selected Exercises Umm
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Selected Exercises Umm
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Selected Exercises Umm
 - Personalized Recommendations
 - Answers To Selected Exercises Umm User Reviews and Ratings
 - Answers To Selected Exercises Umm and Bestseller Lists

5. Accessing Answers To Selected Exercises Umm Free and Paid eBooks
 - Answers To Selected Exercises Umm Public Domain eBooks
 - Answers To Selected Exercises Umm eBook Subscription Services
 - Answers To Selected Exercises Umm Budget-Friendly Options
6. Navigating Answers To Selected Exercises Umm eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers To Selected Exercises Umm Compatibility with Devices
 - Answers To Selected Exercises Umm Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Selected Exercises Umm
 - Highlighting and Note-Taking Answers To Selected Exercises Umm
 - Interactive Elements Answers To Selected Exercises Umm
8. Staying Engaged with Answers To Selected Exercises Umm
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Selected Exercises Umm
9. Balancing eBooks and Physical Books Answers To Selected Exercises Umm
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Selected Exercises Umm
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers To Selected Exercises Umm
 - Setting Reading Goals Answers To Selected Exercises Umm
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers To Selected Exercises Umm
 - Fact-Checking eBook Content of Answers To Selected Exercises Umm
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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