

Types of Chemical Reaction Worksheet

Balance the reactions **1 to 6** and indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented:

- | | |
|---|-------------------------------------|
| 1. $2 \text{NaBr} + \text{Ca(OH)}_2 \rightarrow \text{CaBr}_2 + 2 \text{NaOH}$ | Reaction Type : single displacement |
| 2. $2 \text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow 3 (\text{NH}_4)_2\text{SO}_4$ | Reaction Type : synthesis |
| 3. $4 \text{C}_3\text{H}_8\text{O} + 27 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$ | Reaction Type : combustion |
| 4. $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$ | Reaction Type : single displacement |
| 5. $\text{Li}_3\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$ | Reaction Type : double displacement |
| 6. $\text{HBr} + 3 \text{Al(OH)}_3 \rightarrow \text{H}_2\text{O} + 3 \text{AlBr}_3$ | Reaction Type : double displacement |

Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented in 7 to 20.

- | | |
|--|--|
| 7. $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ | Reaction Type double displacement |
| 8. $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ | Reaction Type double displacement |
| 9. $\text{C}_8\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$ | Reaction Type combustion |
| 10. $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ | Reaction Type single displacement |
| 11. $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ | Reaction Type decomposition |
| 12. $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$ | Reaction Type synthesis |
| 13. $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + 2 \text{RbF}$ | Reaction Type double displacement |
| 14. $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + 2 \text{Ag}$ | Reaction Type single displacement |
| 15. $\text{C}_3\text{H}_8\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ | Reaction Type combustion |
| 16. $2 \text{C}_3\text{H}_6 + \text{Fe} \rightarrow \text{Fe(C}_3\text{H}_5)_2$ | Reaction Type synthesis |
| 17. $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ | Reaction Type single displacement |
| 18. $2 \text{MgI}_2 + \text{Mn(SO}_4)_2 \rightarrow 2 \text{MgSO}_4 + \text{MnI}_4$ | Reaction Type double displacement |
| 19. $\text{O}_3 \rightarrow \text{O} + \text{O}_2$ | Reaction Type decomposition |
| 20. $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$ | Reaction Type decomposition |

Answers For Chemical Reactions

**Barron's Educational Series, Albert
Tarendash**



Answers For Chemical Reactions:

Chemical Reactions, Materials and Particles Gill Murphy, 2006-03 Specials are a teaching resource designed for KS3 students whose literacy skills are considerably lower than their age. These books have an older format to counteract the simple text and cover various topics. They include activities, visuals, and assessment sheets as well as teacher pages.

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Wyatt, 1996-06-27 Covers both molecular and reaction dynamics The work presents important theoretical and computational approaches to the study of energy transfer within and between molecules discussing the application of these approaches to problems of experimental interest It also describes time dependent and time independent methods variational and perturbative techniques iterative and direct approaches and methods based upon the use of physical grids of finite sets of

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