Ionic and Covalent Naming Practice Answers

For each of the following questions, determine whether the compound is ionic or covalent and name it appropriately.

1)	Na ₂ CO ₃	sodium carbonate	
2)	P4O10	tetraphosphorus decaoxide	
3)	N ₂ O ₅	dinitrogen pentoxide	
4)	FeSO ₄	iron (II) sulfate	
5)	SiO ₂	silicon dioxide	
6)	FeCl ₃	iron (III) chloride	
7)	CoBr ₂	cobalt (II) bromide	
8)	B ₂ H ₄	diboron tetrahydride	
9)	co	carbon monoxide	
10)	PCI ₃	phosphorus trichloride	

For each of the following questions, determine whether the compound is ionic or covalent and write the appropriate formula for it.

11)	dinitrogen trioxide	N2O3
12)	nitrogen dioxide	NO ₂
13)	sulfur pentoxide	SO _s
14)	lithium acetate	LiC ₂ H ₃ O ₂
15)	phosphorus trifluoride	PF ₃
16)	vanadium (V) oxide	V ₂ O ₅
17)	aluminum hydroxide	AI(OH)3
18)	zinc sulfide	ZnS
19)	silicon tetrafluoride	SiF4
20)	silver phosphate	Ag ₃ PO ₄

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Sebastian Brünink

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The Covalent Bond Henry Sinclair Pickering,1977 Polar Covalence R Sanderson,2012-12-02 Polar Covalence provides a detailed account of a successful approach to understanding chemistry from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic properties This text then describes chemical bonding which can only occur when the nuclei of both atoms can attract the same electrons Other chapters consider the bond energy of multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the relative roles of principles and practice in the

teaching of inorganic and general chemistry This book is a valuable resource for chemists and students

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Kok Koon Leong,2021-05-03 This book introduces the principles behind chemical bonding to teenagers between the ages of fifteen to seventeen Topics covered include ionic bonding covalent bonding and metallic bonding

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