

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Answers To Chapter Test B Chemical Bonding

Minjie Lin



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Table of Contents Answers To Chapter Test B Chemical Bonding

1. Understanding the eBook Answers To Chapter Test B Chemical Bonding
 - The Rise of Digital Reading Answers To Chapter Test B Chemical Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Chapter Test B Chemical Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Chapter Test B Chemical Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Chapter Test B Chemical Bonding

- Personalized Recommendations
 - Answers To Chapter Test B Chemical Bonding User Reviews and Ratings
 - Answers To Chapter Test B Chemical Bonding and Bestseller Lists
5. Accessing Answers To Chapter Test B Chemical Bonding Free and Paid eBooks
 - Answers To Chapter Test B Chemical Bonding Public Domain eBooks
 - Answers To Chapter Test B Chemical Bonding eBook Subscription Services
 - Answers To Chapter Test B Chemical Bonding Budget-Friendly Options
 6. Navigating Answers To Chapter Test B Chemical Bonding eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers To Chapter Test B Chemical Bonding Compatibility with Devices
 - Answers To Chapter Test B Chemical Bonding Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Chapter Test B Chemical Bonding
 - Highlighting and Note-Taking Answers To Chapter Test B Chemical Bonding
 - Interactive Elements Answers To Chapter Test B Chemical Bonding
 8. Staying Engaged with Answers To Chapter Test B Chemical Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Chapter Test B Chemical Bonding
 9. Balancing eBooks and Physical Books Answers To Chapter Test B Chemical Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Chapter Test B Chemical Bonding
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answers To Chapter Test B Chemical Bonding
 - Setting Reading Goals Answers To Chapter Test B Chemical Bonding
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answers To Chapter Test B Chemical Bonding

- Fact-Checking eBook Content of Answers To Chapter Test B Chemical Bonding
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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